

Nuclear Decay

Types of Radiation

1. Complete the table below to compare the properties of alpha, beta, and gamma radiation.

Name	Alpha	Beta	Gamma
Greek letter	α	β -	γ
Symbol	${}^4_2\text{He}$	${}^0_{-1}e$	γ
Composition	2 protons, 2 neutrons	an electron	radiation (light)
Charge	0	-1	0
Stopped by	paper, skin, clothing	aluminum	lead

Nuclear Decay Reactions

In problems 2 – 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.

- _____ 2. ${}^{14}_6\text{C} \rightarrow {}^{14}_7\text{N} + \text{_____}$ beta A. alpha emission
- _____ 4. ${}^{218}_{84}\text{Po} \rightarrow {}^{214}_{82}\text{Pb} + \text{_____}$ alpha B. beta emission
- _____ 5. $\text{_____} + {}^0_{-1}e \rightarrow {}^{147}_{60}\text{Nd}$ electron capture C. electron capture

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

6. Decay of radium-223 by alpha (α) emission.



7. Decay of iodine-126 by electron (β) emission.



8. How are the atomic number and mass number of a nuclide affected by Beta Decay?

The mass number does not change but the atomic number increases by 1.

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

The atomic number and mass number do not change.

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why?

It depends on the type of reaction. A fusion reaction will have greater mass in the products because it is combining two nuclei. A fission reaction will not change mass.

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JL Elias

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Table of Contents Math Skills Nuclear Decay Answer Key

1. Understanding the eBook Math Skills Nuclear Decay Answer Key
 - The Rise of Digital Reading Math Skills Nuclear Decay Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Math Skills Nuclear Decay Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Math Skills Nuclear Decay Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Math Skills Nuclear Decay Answer Key
 - Personalized Recommendations
 - Math Skills Nuclear Decay Answer Key User Reviews and Ratings
 - Math Skills Nuclear Decay Answer Key and Bestseller Lists
5. Accessing Math Skills Nuclear Decay Answer Key Free and Paid eBooks
 - Math Skills Nuclear Decay Answer Key Public Domain eBooks

- Math Skills Nuclear Decay Answer Key eBook Subscription Services
- Math Skills Nuclear Decay Answer Key Budget-Friendly Options
- 6. Navigating Math Skills Nuclear Decay Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Math Skills Nuclear Decay Answer Key Compatibility with Devices
 - Math Skills Nuclear Decay Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Math Skills Nuclear Decay Answer Key
 - Highlighting and Note-Taking Math Skills Nuclear Decay Answer Key
 - Interactive Elements Math Skills Nuclear Decay Answer Key
- 8. Staying Engaged with Math Skills Nuclear Decay Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Math Skills Nuclear Decay Answer Key
- 9. Balancing eBooks and Physical Books Math Skills Nuclear Decay Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Math Skills Nuclear Decay Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Math Skills Nuclear Decay Answer Key
 - Setting Reading Goals Math Skills Nuclear Decay Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Math Skills Nuclear Decay Answer Key
 - Fact-Checking eBook Content of Math Skills Nuclear Decay Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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