

| Question | | | Answer | Marks | Guidance |
|----------|-----|-------|---|-------|--|
| 2 | (a) | (i) | Al^{3+} ✓ SO_4^{2-} ✓ | 2 | |
| | | (ii) | $Al_2O_3(s) + 3H_2SO_4(aq) \rightarrow Al_2(SO_4)_3(aq) + 3H_2O(l)$ Correct species AND correctly balanced ✓ state symbols on correct species ✓ | 2 | ALLOW multiples |
| | | (iii) | (The number of) water(s) of crystallisation ✓ | 1 | IGNORE hydrated OR hydrous OR 'contains water' |
| | | (iv) | First check the answer on the answer line. If answer = 16, award 3 marks Correctly calculates amount of $Al_2(SO_4)_3$: $6.846 / 342.3 = 0.02(00)$ mol ✓ Correctly calculates amount of H_2O : $5.760 / 18.0 = 0.32(0)$ mol ✓ Correctly calculates whole number ratio of mol of H_2O : $Al_2(SO_4)_3$ to give $x = 16$ ✓ | 3 | If there is an alternative answer, check to see if there is any ECF credit possible using working below ALLOW as ECF from $12.606/342.3 = 0.0368(273)$ AND $0.32/0.0368(273)$ To give $x = 9$ for two marks ALLOW calculator value or rounding to 2 significant figures or more BUT IGNORE 'trailing' zeroes, eg 0.200 allowed as 0.2. ALLOW ECF for calculation of correctly rounded whole number value of H_2O from incorrect mol of H_2O and / or incorrect mol of $Al_2(SO_4)_3$ BUT x must be a whole number ALLOW alternative method Mol of $Al_2(SO_4)_3$: $6.846 / 342.3 = 0.02(00)$ mol (first mark) Molar mass of $Al_2(SO_4)_3 \cdot xH_2O$: $12.606 / 0.02(00) = 630.3$ g mol ⁻¹ (second mark) Mass of water per mol = $630.3 - 342.3 = 288$ AND $288/18$ to give $x = 16$ (third mark) |

Mark Scheme June 2013 Core One Ocr

Lauren Gardner



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Table of Contents Mark Scheme June 2013 Core One Ocr

1. Understanding the eBook Mark Scheme June 2013 Core One Ocr
 - The Rise of Digital Reading Mark Scheme June 2013 Core One Ocr
 - Advantages of eBooks Over Traditional Books
2. Identifying Mark Scheme June 2013 Core One Ocr
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mark Scheme June 2013 Core One Ocr
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mark Scheme June 2013 Core One Ocr
 - Personalized Recommendations
 - Mark Scheme June 2013 Core One Ocr User Reviews and Ratings
 - Mark Scheme June 2013 Core One Ocr and Bestseller Lists
5. Accessing Mark Scheme June 2013 Core One Ocr Free and Paid eBooks
 - Mark Scheme June 2013 Core One Ocr Public Domain eBooks
 - Mark Scheme June 2013 Core One Ocr eBook Subscription Services

- Mark Scheme June 2013 Core One Ocr Budget-Friendly Options
- 6. Navigating Mark Scheme June 2013 Core One Ocr eBook Formats
 - ePub, PDF, MOBI, and More
 - Mark Scheme June 2013 Core One Ocr Compatibility with Devices
 - Mark Scheme June 2013 Core One Ocr Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mark Scheme June 2013 Core One Ocr
 - Highlighting and Note-Taking Mark Scheme June 2013 Core One Ocr
 - Interactive Elements Mark Scheme June 2013 Core One Ocr
- 8. Staying Engaged with Mark Scheme June 2013 Core One Ocr
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mark Scheme June 2013 Core One Ocr
- 9. Balancing eBooks and Physical Books Mark Scheme June 2013 Core One Ocr
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mark Scheme June 2013 Core One Ocr
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mark Scheme June 2013 Core One Ocr
 - Setting Reading Goals Mark Scheme June 2013 Core One Ocr
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mark Scheme June 2013 Core One Ocr
 - Fact-Checking eBook Content of Mark Scheme June 2013 Core One Ocr
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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