



# Group 18: Noble Gases

18
He 2 1-18
18
Ne 10 1-18
Ar 18 1-18
Kr 36 1-18
Xe 54 1-18
Rn 86 1-18

- Are completely nonreactive since they have eight valence electrons, making a stable octet.
- Kr and Xe can be forced, in the laboratory, to give up some valence electrons to react with fluorine.
- Since noble gases do not naturally bond to any other elements, one atom of noble gas is considered to be a molecule of noble gas. This is called a monatomic molecule. Ne represents an atom of Ne and a molecule of Ne.

# Mark Rosengarten Labs

**Jianjun Gao**



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