

MATH SKILLS TRANSPARENCY WORKSHEET**2**

Visualizing the Conservation of Mass

**Use with Chapter 3,
Section 3.2**

1. How many potassium atoms are in the reactants of the chemical reaction shown? How many are in the products? _____
2. How many oxygen atoms are in the reactants of the chemical reaction shown? How many are in the products? _____
3. How many hydrogen atoms are in the reactants of the chemical reaction shown? How many are in the products? _____
4. Assume that the chemical reaction shown started out having a total of 15 g of potassium and water. How much potassium hydroxide and hydrogen gas will be produced by the chemical reaction? Show your work.

5. Assume that the chemical reaction shown started out having 6 atoms of potassium and 6 molecules of water. How many molecules of potassium hydroxide will be produced by the chemical reaction? How many hydrogen atoms will result?

6. Assume that you are working with a chemical reaction that decomposes water into hydrogen and oxygen. You begin with 36 grams of water and end with 32 grams of oxygen. If all of the water decomposes, how many grams of hydrogen gas will result?

7. Assume that you are working with a chemical reaction that synthesizes salt from sodium and chlorine. You begin with 70.9 g of chlorine. You synthesize 116.90 grams of salt. If all of the reactants were used up, how many grams of sodium did you begin with?

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**Holt Rinehart & Winston, Holt,
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