

MARKING SCHEME

TABLE 1

Experiment number	I	II	III
Final burette reading (cm ³)	13.0	25.9	38.7
Initial burette reading (cm ³)	0.0	13.0	25.9
Volume of solution B used (cm ³)	13.0	12.9	12.8

Complete table	1
Decimal place	1
Accuracy (tied to school value	1
Arithmetic (correct)	1

- (a) Average volume of solution B used

$$\frac{13.0 + 12.9 + 12.8}{3} \sqrt{\frac{1}{2}} = 12.9 \text{ cm}^3 \sqrt{\frac{1}{2}}$$
- (b) Molarity of solution A

$$\text{Molarity} = \frac{\text{mass}}{\text{RMM}} \div \text{litre} = \frac{4.0}{40} = 0.1 \text{ M} \sqrt{1}$$
- (c) Molarity of solution B

$$\text{HCl (aq)} + \text{NaOH (aq)} \rightarrow \text{NaCl (aq)} + \text{H}_2\text{O (l)}$$

 Mole ratio 1 : 1
 Moles of NaOH = $0.1 \times \frac{25}{1000} \sqrt{\frac{1}{2}}$

$$= 0.0025 \text{ moles} \sqrt{\frac{1}{2}}$$

 Moles of HCl = moles of NaOH = 0.0025 moles
 Molarity of HCl = $\frac{0.0025 \times 1,000}{12.9} \sqrt{\frac{1}{2}}$

$$= 0.19379 \text{ M} \sqrt{\frac{1}{2}}$$

TABLE 2

Experiment number	I	II	III
Final burette reading (cm ³)	5.0	10.0	15.0
Initial burette reading (cm ³)	0.0	5.0	10.0
Volume of solution B used (cm ³)	5.0	5.0	5.0

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