

## 4.7 CHEMISTRY (12%)

### 4.7.1 Chemistry Paper 1 (12%)

4. (a) Ether (1 mark)
- (b) Add water to each compound.  
Add  $\text{Na}_2\text{CO}_3$  or  $\text{NaHCO}_3$  to a suitable part of each sample.  
 $\text{C}_2\text{H}_5\text{CO}$  no effervescence (1 mark)  
 $\text{C}_2\text{H}_5\text{CO}_2$  effervescence (1 mark)  
or add acidified potassium dichromate (VII) and warm  
 $\text{C}_2\text{H}_5\text{CO}$  colourless change to green  
 $\text{C}_2\text{H}_5\text{CO}_2$  no change  
Or add acidified Potassium Manganate (VII)  
 $\text{C}_2\text{H}_5\text{CO}$  + will be decolourised  
 $\text{C}_2\text{H}_5\text{CO}_2$  no change (1 mark)
5. (a) Ether ( $\text{Na}_2\text{CO}_3$ ) (1 mark)
- (b) — Sodium is very reactive (more electropositive) (1 mark)  
— More reactive than carbon
6. (a) Ether  
Sodium lamp, used in nuclear reactions  
Sodium vapour, sodium amalgams  
( $\text{Na}_2\text{CO}_3$ ), Electrodes and cathodes, etc. (1 mark)
7. (a) Endothermic change, which one mole of crystal lattice is broken into its ions in gaseous state. (1 mark)
- (b) Endothermic reaction (percent) (1 mark)
8. (a) Boyle's law: The volume of a fixed mass of a gas is inversely proportional to its pressure or the product of pressure and volume is constant at a fixed temperature. (1 mark)

# Mark Schemes Chemistry Edexcel 2015

**Iseb Iseb**



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