

HEAT AND ITS MEASUREMENT

18/07/20

Heat for energy transfer has measured in units of calories or joules. When measuring temperature change, the heat transferred can be calculated using the formula:

$$Q = mc\Delta T$$

(Q is energy, m is mass, c is specific heat capacity, and ΔT is the temperature change)

During a phase change, we use the formula:

$$Q = mL$$

(Q is energy, L is latent heat, and m is mass)

Solve the following problems.

1. How many joules of heat are given off when 5.0 g of water cools from 100°C to 25°C? (Specific heat of water = 4.18 J/g°C)	- 1045 J
2. How many calories are given off by the water in Problem 1?	- 25 cal
3. How many joules does it take to heat 20 g of ice from 0°C to 25°C? (Specific heat of water = 4.18 J/g°C)	835 J
4. How many calories are given off when 50 g of water cools from 100°C to 25°C? (Specific heat of water = 4.18 J/g°C)	4975 cal
5. How many joules of heat are necessary to raise the temperature of 25 g of water from 10°C to 45°C?	3715 J
6. How many calories are given off when 50 g of water (100°C) freezes? (Heat of fusion = 79.7 cal/g)	3985 cal

Het And Its Measurement Answer Key

David Baud



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