

Percent Composition and Molecular Formula Worksheet

1. Calculate the percent of nitrogen in urea, NH_2CONH_2 .
2. Calculate the percentage of water in zinc sulfate heptahydrate.
3. Calculate the percentage of oxygen in potassium chlorate.
4. Calculate the percentage of tin in potassium stannate trihydrate, $\text{K}_2\text{SnO}_3 \cdot 2\text{H}_2\text{O}$

Write the molecular formulas of the following compounds:

5. A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole.
6. A compound with an empirical formula of $\text{C}_4\text{H}_4\text{O}$ and a molar mass of 136 grams per mole.
7. A compound with an empirical formula of CFBrO and a molar mass of 254.7 grams per mole.
8. A compound with an empirical formula of $\text{C}_2\text{H}_5\text{N}$ and a molar mass of 46 grams per mole.

Answer the following questions:

9. The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid.
10. The molar mass for question #9 was determined by experiment to be 60.0 g/mol. What is the molecular formula?
11. What's the empirical formula of a molecule containing 18.7% lithium, 16.3% carbon, and 65.0% oxygen?
12. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound?
13. When 2.5000 g of an oxide of mercury, (Hg_xO_y) is decomposed into the elements by heating, 2.405 g of mercury are produced. Calculate the empirical formula.
14. The compound benzamide has the following percent composition. What is the empirical formula?
C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 %
15. A component of protein called serine has an approximate molar mass of 100 g/mole. If the percent composition is as follows, what is the empirical and molecular formula of serine?
C = 34.95 % H= 6.844 % O = 46.56 % N= 13.59 %
16. Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields CO_2 , H_2O , and N_2 as products. If the combustion of 9.71 g of aniline yields 6.63 g H_2O and 1.46 g N_2 , what is its empirical formula?
17. The molar mass of aniline is 93 g/mol. What is its molecular formula?

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Table of Contents Formula Writing Cavalcade

1. Understanding the eBook Formula Writing Cavalcade
 - The Rise of Digital Reading Formula Writing Cavalcade
 - Advantages of eBooks Over Traditional Books
2. Identifying Formula Writing Cavalcade
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Formula Writing Cavalcade
 - User-Friendly Interface
4. Exploring eBook Recommendations from Formula Writing Cavalcade

- Personalized Recommendations
- Formula Writing Cavalcade User Reviews and Ratings
- Formula Writing Cavalcade and Bestseller Lists
- 5. Accessing Formula Writing Cavalcade Free and Paid eBooks
 - Formula Writing Cavalcade Public Domain eBooks
 - Formula Writing Cavalcade eBook Subscription Services
 - Formula Writing Cavalcade Budget-Friendly Options
- 6. Navigating Formula Writing Cavalcade eBook Formats
 - ePub, PDF, MOBI, and More
 - Formula Writing Cavalcade Compatibility with Devices
 - Formula Writing Cavalcade Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Formula Writing Cavalcade
 - Highlighting and Note-Taking Formula Writing Cavalcade
 - Interactive Elements Formula Writing Cavalcade
- 8. Staying Engaged with Formula Writing Cavalcade
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Formula Writing Cavalcade
- 9. Balancing eBooks and Physical Books Formula Writing Cavalcade
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Formula Writing Cavalcade
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Formula Writing Cavalcade
 - Setting Reading Goals Formula Writing Cavalcade
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Formula Writing Cavalcade

- Fact-Checking eBook Content of Formula Writing Cavalcade
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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