

Chemistry – Unit 7: Chemical Reactions **Worksheet 1: Reacting Atoms**

Background

Though they often go unnoticed, several factors make up a chemical change. Three are highlighted in each of the four reactions.

Chemical Change: Chemical Model of Mass

Conservation of Mass

Chemical Formula

Subscripts vs. Coefficients

Question: What effect do the coefficients of a coefficient on reactants?

Procedures

1. Two types of chemical reactions were the primary focus of this experiment. The first was a synthesis reaction. The second was a decomposition reaction. The first reaction was the reaction of hydrogen and oxygen to form water. The second reaction was the reaction of hydrogen peroxide to form water and oxygen. The first reaction was a synthesis reaction. The second was a decomposition reaction. The first reaction was the reaction of hydrogen and oxygen to form water. The second reaction was the reaction of hydrogen peroxide to form water and oxygen.
2. The first reaction was a synthesis reaction. The second reaction was a decomposition reaction. The first reaction was the reaction of hydrogen and oxygen to form water. The second reaction was the reaction of hydrogen peroxide to form water and oxygen.
3. The first reaction was a synthesis reaction. The second reaction was a decomposition reaction. The first reaction was the reaction of hydrogen and oxygen to form water. The second reaction was the reaction of hydrogen peroxide to form water and oxygen.
4. The first reaction was a synthesis reaction. The second reaction was a decomposition reaction. The first reaction was the reaction of hydrogen and oxygen to form water. The second reaction was the reaction of hydrogen peroxide to form water and oxygen.

Modeling Chemistry Unit 7 Rearranging Atoms Key

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