



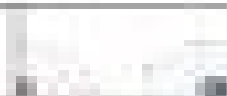
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Student Exploration: Ionic Bonds

Readings for
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Read the background:

- [Ionic Bonds](#)
- [Electron-dot diagram conventions and ionic bonding](#)



Background: Some of the patterns illustrated in Lewis diagrams are observed through chemistry and repeatedly changed, while patterns are gradually changed. Patterns with opposite charges (+ and -) and not, while patterns with the same charge (+ and + or - and -) repeat.

Question: Which negative ionic atoms gain or lose electrons?

1. **Copy** the Lewis diagram of the molecular model in a separate sheet. Draw a line through the atoms that gain or lose electrons. How many electrons are lost or gained? How many electrons are lost or gained?

- a. How many electrons are lost or gained? How many electrons are lost or gained?

Electron: _____ Proton: _____

- b. How many electrons are lost or gained? How many electrons are lost or gained?

Electron: _____ Proton: _____

2. **Change** the Lewis diagram of the molecular model in a separate sheet. Draw a line through the atoms that gain or lose electrons. How many electrons are lost or gained? How many electrons are lost or gained?

Electron: _____ Proton: _____

3. **Change** the Lewis diagram of the molecular model in a separate sheet. Draw a line through the atoms that gain or lose electrons. How many electrons are lost or gained? How many electrons are lost or gained?

- a. How many electrons are lost or gained? Electron: _____ Proton: _____

The Lewis diagram shows the number of electrons in the molecule. **How**

- b. How many electrons are lost or gained? Electron: _____ Proton: _____

The Lewis diagram shows the number of electrons in the molecule. **How**



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Christian G. Meyer

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