

7&8 • Atomic Structure & Periodicity

PRACTICE TEST

$$Rhc = A$$

$$A = 2.18 \times 10^{-18} \text{ J}$$

$$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$$

$$R = 1.097 \times 10^7 \text{ m}^{-1}$$

$$c = 3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}$$

$$\text{mass of an electron} = 9.11 \times 10^{-31} \text{ kg}$$

1. What wavelength corresponds to a frequency of $8.22 \times 10^9 \text{ Hz}$?

$$\lambda = \frac{c}{\nu}$$

- a) 0.307 m d) 0.110 m

b) 0.0365 m $\frac{3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}{8.22 \times 10^9 \text{ s}^{-1}} = 0.036496 \text{ m}$

c) 0.122 m

2. A radio station transmits at 110 MHz ($110 \times 10^6 \text{ Hz}$). What wavelength is this radio wave?

- a) $3.65 \times 10^{-5} \text{ m}$ c) $3.81 \times 10^{-5} \text{ m}$

- b) 3.30 m d) 2.73 m

$$\lambda = \frac{c}{\nu} = \frac{3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}}{110 \times 10^6 \text{ s}^{-1}} = 2.73 \text{ m}$$

3. Which one of the following is NOT a proper unit for frequency?

- a) Hz

- b) s^{-1}

- c) $\text{m}\cdot\text{s}^{-1}$

- d) $\frac{1}{\text{sec}}$

velocity

4. Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen spectrum.

$$n=6 \rightarrow n=2$$

$$E_6 = -2.18 \times 10^{-18} \text{ J} - E_2 = -2.18 \times 10^{-18} \text{ J}$$

- a) 0.12334 m

- d) $4.1029 \times 10^{-7} \text{ m}$

- b) 24.373 m

- e) 36.559 m

- c) $2.7353 \times 10^{-7} \text{ m}$

$$E = \frac{hc}{\lambda} \quad \lambda = \frac{hc}{E} = 4.1 \times 10^{-7} \text{ m}$$

$$\Delta E = 4.84 \times 10^{-19} \text{ J}$$

5. What is the relationship between the energy of a photon of light and its frequency?

- a) $E = \nu$

- d) $E = \frac{1}{h\nu}$

- b) $E = \frac{h}{\nu}$

- e) $E = \frac{\nu}{h}$

- c) $E = h\nu$

6. What is the energy needed to raise an electron in the hydrogen atom from the second energy level to the third energy level?

$$\Delta E = E_3 - E_2$$

- a) $1.52 \times 10^{-19} \text{ J}$

- d) $4.48 \times 10^{-19} \text{ J}$

- b) $3.63 \times 10^{-19} \text{ J}$

- e) $3.03 \times 10^{-19} \text{ J}$

- c) $2.18 \times 10^{-19} \text{ J}$

$$-2.18 \times 10^{-18} \text{ J} - (-2.18 \times 10^{-18} \text{ J}) = 3.03 \times 10^{-19} \text{ J}$$

7. What is the de Broglie wavelength of an electron moving at 80.0% the speed of light.

- a) $3.03 \times 10^{-12} \text{ m}$

- c) $3.30 \times 10^{11} \text{ m}$

- b) $2.42 \times 10^{-12} \text{ m}$

- d) $1.59 \times 10^{-25} \text{ m}$

$$\lambda = \frac{h}{mv} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{9.11 \times 10^{-31} \text{ kg} \cdot 0.8 \cdot 3.00 \times 10^8 \text{ m}\cdot\text{s}^{-1}} = 3.03 \times 10^{-12} \text{ m}$$

8. What resultant is expected from the interference of the two waves shown below?



- a)

- c)

- b)

- d)

9. Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?

- a) n

- c) m_l

- b) ℓ

- d) m_s

Electron Configuration Chem Worksheet 5 6 Key

Christian G. Meyer



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