

The Hong Kong University of Science and Technology
Department of Chemistry

CHEM 1020 General Chemistry III
Homework (Ch4 + Ch5.1-VSEPR Theory)

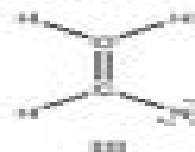
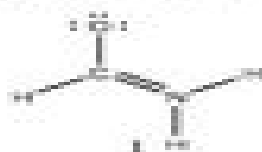
Fall 2016-2017

Some questions are extracted from "Chemistry: An Atomic First Approach (2nd Edition)"

5. For each multiple choice questions, select the most appropriate answer.

- i) Which of the following statements regarding ionic bonding is FALSE?
- a) Ionic bonding is the electrostatic interaction between positive and negative charged ions.
 - b) Lattice energy can give reference to the bond strength of ionic bonding.
 - c) Sharing of an electron pair gives an ionic bond.
 - d) Larger charge of ions gives stronger ionic bonding.

- ii) Which of the following is NOT a possible Lewis structure of formamide ($\text{HCO(NH}_2\text{)}$)?



- a) I & II
- b) III & IV
- c) II, III & IV
- d) All of them

- iii) Which of the following statements regarding formal charge is CORRECT?

- a) The sum of the formal charges of an ion cannot be zero.
- b) The most stable Lewis structure(s) of a species must have zero formal charge for every atom.
- c) The formal charge of carbon in carbon monoxide (CO) is the same as in formaldehyde ($\text{H}_2\text{C=O}$).
- d) The formal charge of sodium in the ionic compound sodium hydroxide (NaOH) is zero.

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