

JELET-2023
Paper-II (For Pharmacy Candidates)

(Booklet Number)

Duration: 2 Hours

No. of MCQ : 100

Full Marks: 100

INSTRUCTIONS

1. All questions are of objective type having four answer options for each.
2. Category-I MCQ : Carry 1 mark each and only one option is correct. In case of incorrect answer or any combination of more than one answer, $\frac{1}{4}$ mark will be deducted.
3. Questions must be answered on OMR sheet by darkening the appropriate bubble marked A, B, C, or D.
4. Use only **Black/Blue ink ball point pen** to mark the answer by complete filling up of the respective bubbles.
5. Write question booklet number and your roll number carefully in the specified locations of the **OMR Sheet**. Also fill appropriate bubbles.
6. Write your name (in block letter), name of the examination center and put your signature (as appeared in Admit Card) in appropriate boxes in the OMR Sheet.
7. The OMR Sheet is liable to become invalid if there is any mistake in filling the correct bubbles for question booklet number/roll number or if there is any discrepancy in the name / signature of the candidate, name of the examination center. The OMR Sheet may also become invalid due to folding or putting stray marks on it or any damage to it. The consequence of such invalidation due to incorrect marking or careless handling by the candidate will be sole responsibility of candidate.
8. Candidates are not allowed to carry any written or printed material, calculator, pen, log-table, wristwatch, any communication device like mobile phones, bluetooth etc. inside the examination hall. Any candidate found with such prohibited items will be **reported against** and his/her candidature will be summarily cancelled.
9. Rough work must be done on the question booklet itself. Additional blank pages are given in the question booklet for rough work.
10. Hand over the OMR Sheet to the invigilator before leaving the Examination Hall.
11. Candidates are allowed to take the Question Booklet after Examination is over.

Signature of the Candidate : _____
(as in Admit Card)

Signature of the Invigilator : _____

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web 1 molarity mv mass molar mass $\times 0.4500 \text{ L} = 0.825 \text{ g}$ 141 9579 g/mol $\times 0.0129 \text{ M} = 2 \text{ molality}$ 0.825 g 141 9579 g/mol 0.00581158 mol 0.00581158 mol 0.4500 kg 0.0129 M 3 mole fraction $n_2 = \frac{0.825 \text{ g}}{141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ $n_1 = \frac{450.0 \text{ g}}{18.015 \text{ g/mol}} = 24.97918401 \text{ mol}$

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web apr 7 2014 molarity moles of solute liters of solution 8.4 2.2 a first convert 250 ml to liters 250 1000 = 0.25 then calculate molarity 5 moles / 0.25 liters = 20 M 3 c a solution with molarity 2 requires 2 M of NaOH per liter so 4 x 2.8 M = 11.2 M a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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molarity of a solution containing 325 g of nacl dissolved in 750 ml of solution 3 140 g of kcl is dissolved in 600 ml of water what is the molarity 4 724 4 g of ammonium phosphate in 4500 ml of alcohol

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