

Types of Chemical Reactions Lab

Chemical reactions can be categorized into five different types of reactions:

Synthesis

Decomposition

Single Displacement

Double Displacement

Combustion

In this laboratory activity, you will experiment with various chemical reactions and **classify them** based on their **reactants** and **products**.

Materials

- Copper pellets (Cu)
- Magnesium strips (Mg)
- Zinc metal (Zn)
- Silver nitrate (AgNO_3)
- Copper (II) sulfate (CuSO_4)
- 1M Hydrochloric acid (HCl)
- Hydrogen peroxide (H_2O_2)
- Potassium iodide (KI)
- Lead (II) nitrate (PbNO_3)
- Test tubes (3)
- Test tube holder
- Bunsen burner

Procedure

Reaction A

1. Obtain a small number of copper (Cu) pellets. Place them into your test tube.
2. Carefully pour (or drip) silver nitrate (AgNO_3) into the test tube until it covers the copper.
3. Allow the reaction in the test tube to proceed for several minutes. Observe any changes.

Reaction B

4. Obtain a small chunk of zinc (Zn) metal. Place it into your test tube.
5. Carefully pour (or drip) hydrochloric acid (HCl) into the test tube until it covers the zinc.
6. Allow the reaction in the test tube to proceed for several minutes. Observe any changes.

Reaction C

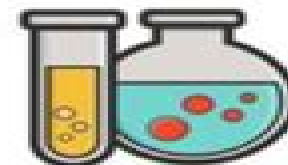
7. Add approximately 10 drops of lead nitrate (PbNO_3) into a small test tube.
8. Add 2 drops of potassium iodide (KI) to the test tube. Observe any changes.

Reaction D

9. Obtain a few chunks of copper (II) sulfate. Place them into your test tube.
10. Light your Bunsen burner. Using a test tube holder, heat the test tube of copper (II) sulfate in the Bunsen burner flame for roughly 2 minutes. Observe any changes.

Reaction E

11. Obtain a strip of magnesium. Carefully hold the magnesium in the Bunsen burner flame using tongs. When the reaction starts, *do not look directly at the magnesium*. Observe any changes.



Experiment 14 Types Of Chemical Reactions Key

M Mark



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