B

## ELECTRONS IN ATOMS

Conceptsed Curriculum Standar
Concrete concepts

More abstract concepts or math/problems solving

Standard Curriculum
Core content

Extension topics

Honors Curriculum
Core honors content
Opcious to accelerate

#### SECTION 13.1 MODELS OF THE ATOM (pages 361-366)

This section summarizes the development of atomic theory. It also explains the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.

#### ➤ The Evolution of Atomic Models (pages 361-363)

1. What are the chemical properties of atoms, ions, and molecules related to?

The properties are related to the arrangement of electrons.

Complete the table about atomic models and the scientists who developed them.

Scientist	Model of Atom
Dalton	The atom is a solid indivisible mass.
Thomson	The atom is a ball of positive charge with electrons stuck into the ball.
Rutherford	Most of an atom's mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty space.
Bohr	Electrons are arranged in concentric circular paths around the nucleus.

- The energy level of an electron is the region around the nucleus where the electron is likely to be moving
- Is the following sentence true or false? The electrons in an atom can exist between energy levels.
- Circle the letter of the term that completes the sentence correctly. A quantum of energy is the amount of energy required to
  - a. move an electron from its present energy level to the next lower one
  - b. maintain an electron in its present energy level
  - c.) move an electron from its present energy level to the next higher one
- 6. In general, the higher the electron is on the energy ladder, the

farther it is from the nucleus.

# **Electrons In Atoms Section 5 2 Key Answers**

**A Loxley** 

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### **Table of Contents Electrons In Atoms Section 5 2 Key Answers**

- 1. Understanding the eBook Electrons In Atoms Section 5 2 Key Answers
  - The Rise of Digital Reading Electrons In Atoms Section 5 2 Key Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Electrons In Atoms Section 5 2 Key Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - $\circ$  Features to Look for in an Electrons In Atoms Section 5 2 Key Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Electrons In Atoms Section 5 2 Key Answers
  - Personalized Recommendations
  - Electrons In Atoms Section 5 2 Key Answers User Reviews and Ratings
  - Electrons In Atoms Section 5 2 Key Answers and Bestseller Lists
- 5. Accessing Electrons In Atoms Section 5 2 Key Answers Free and Paid eBooks
  - Electrons In Atoms Section 5 2 Key Answers Public Domain eBooks
  - Electrons In Atoms Section 5 2 Key Answers eBook Subscription Services
  - Electrons In Atoms Section 5 2 Key Answers Budget-Friendly Options
- 6. Navigating Electrons In Atoms Section 5 2 Key Answers eBook Formats

- o ePub, PDF, MOBI, and More
- Electrons In Atoms Section 5 2 Key Answers Compatibility with Devices
- Electrons In Atoms Section 5 2 Key Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electrons In Atoms Section 5 2 Key Answers
  - Highlighting and Note-Taking Electrons In Atoms Section 5 2 Key Answers
  - Interactive Elements Electrons In Atoms Section 5 2 Key Answers
- 8. Staying Engaged with Electrons In Atoms Section 5 2 Key Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electrons In Atoms Section 5 2 Key Answers
- 9. Balancing eBooks and Physical Books Electrons In Atoms Section 5 2 Key Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electrons In Atoms Section 5 2 Key Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Electrons In Atoms Section 5 2 Key Answers
  - Setting Reading Goals Electrons In Atoms Section 5 2 Key Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electrons In Atoms Section 5 2 Key Answers
  - Fact-Checking eBook Content of Electrons In Atoms Section 5 2 Key Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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