
GCSE Chemistry – Electrolysis Exam Questions and Answers

What is electrolysis? - Answer- passing an electric current through an ionic substance that's molten or in a solution and breaks it down into new substances

What is electrolysis used for? - Answer- To make new substances

What is an electrolyte? - Answer- a liquid that conducts electricity

How are electrolytes made? - Answer- by melting or dissolving ionic compounds

Which part of an electrolyte conducts electricity? - Answer- The free electrons

What is the a positive electrode called? - Answer- Anode

What is a negative electrode called? - Answer- Cathode

What are electrodes made from and why? - Answer- Unreactive metal so they don't take part in the reaction

How do ions become atoms or molecules? - Answer- By losing or gaining electrons

How does an electrolysis circuit become complete? - Answer- A flow of electrons taken from anodes to cathodes.

How do you measure the conductivity of an electrolyte? - Answer- With a conductivity probe

In molten ionic compounds there is only _____ - Answer- one source of ions

Why can molten ionic compounds be electrolysed? - Answer- Because ions can move freely

In the electrolysis of molten lead bromide what is the half equation for the Pb^{2+} ions? - Answer- $\text{Pb}^{2+} + 2\text{e}^{-} \rightarrow \text{Pb}$

In the electrolysis of molten lead bromide what is the half equation for the Br^{-} ions? - Answer- $2\text{Br}^{-} \rightarrow \text{Br}_2 + 2\text{e}^{-}$

Electrolysis Questions And Answers

ML Morrison



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