

# SHORT FORMULA (GYAN SUTRA)

## CHEMISTRY

### PHYSICAL CHEMISTRY

#### ATOMIC STRUCTURE

Estimation of closest distance of approach (derivation) of  $\alpha$ -particle :

$$R = \frac{4\pi KZe^2}{m_{\alpha} V_{\alpha}^2}$$

The radius of a nucleus :

$$R = R_0 (A)^{1/3} \text{ cm}$$

Planck's Quantum Theory :

$$\text{Energy of one photon} = h\nu = \frac{hc}{\lambda}$$

Photoelectric Effect :

$$h\nu = h\nu_0 + \frac{1}{2} m_e v^2$$

Bohr's Model for Hydrogen like atoms :

$$1. \quad mvr = n \frac{h}{2\pi} \quad (\text{Quantization of angular momentum})$$

$$2. \quad E_n = -\frac{E_1}{n^2} \quad Z^2 = 2.178 \times 10^{-18} \frac{Z^2}{n^2} \text{ J/atom} = 13.6 \frac{Z^2}{n^2} \text{ eV} \quad ; \quad E_1 = \frac{-2\pi^2 me^4}{n^2}$$

$$3. \quad r_n = \frac{n^2}{Z} \times \frac{h^2}{4\pi^2 e^2 m} = \frac{0.529 \times n^2}{Z} \text{ \AA}$$

$$4. \quad v = \frac{2\pi ze^2}{nh} = \frac{2.18 \times 10^6 \times z}{n} \text{ m/s}$$

De-Broglie wavelength :

$$\lambda = \frac{h}{mc} = \frac{h}{p} \quad (\text{for photon})$$

Wavelength of emitted photon :

$$\frac{1}{\lambda} = \bar{\nu} = RZ^2 \left( \frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$$

No. of photons emitted by a sample of H atom :

$$\frac{\Delta n (\Delta n + 1)}{2}$$

# Iit Jee Physics Chemistry Maths Formula Booklet

**Career Point Kota**



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