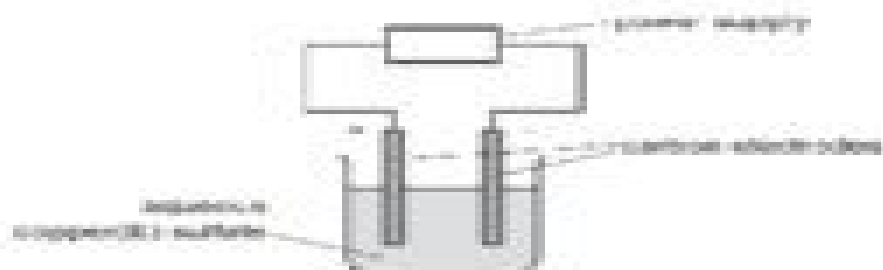


Question 1 * 64/66 (41/44/12/00)

(a) A student electrolyses aqueous copper(II) sulfate using the apparatus shown.



Oxygen gas forms at the positive electrode (anode).

(b) Write an ionic half-equation for the reaction at the negative electrode (cathode). Include state symbols.

_____ [3]

(c) Describe what the student observes at the negative electrode.

_____ [1]

(d) Give two other observations which the student makes during the electrolysis.

1. _____

2. _____

[2]

(e) What difference would the student observe at the positive electrode if the aqueous copper(II) sulfate were replaced by concentrated aqueous copper(II) chloride?

_____ [1]

Question 2 * 64/66 (40/44/12/02/14)

(a) Electrolysis of concentrated aqueous sodium chloride is an important industrial process.

(i) What is meant by the term anode/cathode?

_____ [1]

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Table of Contents Igcse Electrolysis Past Papers

1. Understanding the eBook Igcse Electrolysis Past Papers
 - The Rise of Digital Reading Igcse Electrolysis Past Papers
 - Advantages of eBooks Over Traditional Books
2. Identifying Igcse Electrolysis Past Papers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Igcse Electrolysis Past Papers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Igcse Electrolysis Past Papers
 - Personalized Recommendations
 - Igcse Electrolysis Past Papers User Reviews and Ratings
 - Igcse Electrolysis Past Papers and Bestseller Lists
5. Accessing Igcse Electrolysis Past Papers Free and Paid eBooks
 - Igcse Electrolysis Past Papers Public Domain eBooks
 - Igcse Electrolysis Past Papers eBook Subscription Services
 - Igcse Electrolysis Past Papers Budget-Friendly Options
6. Navigating Igcse Electrolysis Past Papers eBook Formats

- ePub, PDF, MOBI, and More
- Igcse Electrolysis Past Papers Compatibility with Devices
- Igcse Electrolysis Past Papers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Igcse Electrolysis Past Papers
 - Highlighting and Note-Taking Igcse Electrolysis Past Papers
 - Interactive Elements Igcse Electrolysis Past Papers
- 8. Staying Engaged with Igcse Electrolysis Past Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Igcse Electrolysis Past Papers
- 9. Balancing eBooks and Physical Books Igcse Electrolysis Past Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Igcse Electrolysis Past Papers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Igcse Electrolysis Past Papers
 - Setting Reading Goals Igcse Electrolysis Past Papers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Igcse Electrolysis Past Papers
 - Fact-Checking eBook Content of Igcse Electrolysis Past Papers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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