Determine the volume of CO₂ gas given off when excess MgCO₃ is added to 25.0 cm³ of 0.400 mol/dm³ HCl at room temperature and pressure.

Use the following steps.

C

Calculate the number of moles of HCI in 25.0 cm³ of 0.400 mol/dm³ of acid.

$$c = \frac{m}{V}$$
 $m = 0.4 \times 0.025$
 $= 0.01$

Determine the number of moles of CO, gas given off.

Calculate the volume of CO₂ gas given off in cm³.

$$1 \text{ mol} = 24 \text{ dm}^3$$

 $0.005 \text{ mol} = x \text{ dm}^3$
 $x = 0.005 \times 24$
 $= 0.12 \text{ dm}^3$
 $= 120 \text{ cm}^3 (\times 1000)$

1 2 O cm³

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Harvey F Carroll

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