Name_____ Date_____ Per_____

Mole to Grams, Grams to Moles Conversions Worksheet

What are the molecular weights of the following compounds? *(all masses must be to nearest hundredth)*

1) $NaOH$ 2) H_3PO_4

3) H_2O 4) Mn_2Se_7

5) $MgCl_2$ 6) $(NH_4)_2SO_4$

There are three definitions (equalities) of mole. They are:

1 mole = 6.02×10^{23} particles
1 mole = molar mass (could be atomic mass from periodic table or molecular mass)
1 mole = 22.4 L of a gas at STP **(You do not need to worry about this yet)**

Each definition can be written as a set of two conversion factors. They are:

1 mole = molar mass(g) can be written as $\left(\frac{1\text{ mole}}{\text{molar mass (g)}}\right)$ **OR** $\left(\frac{\text{molar mass (g)}}{1\text{ mole}}\right)$

1 mole = 6.02×10^{23} particles can be written as $\left(\frac{1\text{ mole}}{6.02 \times 10^{23}}\right)$ **OR** $\left(\frac{6.02 \times 10^{23}}{1\text{ mole}}\right)$

Solve the following:

1) **How many moles** are in 15 grams of lithium? **(molar mass of lithium is 6.94 g/mole)**

15 ~~grams~~ x $\frac{1\text{ mole}}{6.94\text{~~grams~~}}$ = 2.1614 moles lithium = **2.2 moles Li**

2) **How many grams** are in 2.4 moles of sulfur? **(molar mass of sulfur is 32.07 g/ mole)**

2.4 ~~moles~~ x $\frac{32.07\text{ grams}}{1\text{ ~~mole~~}}$ = 76.97 grams sulfur = **77 g Sulfur**

3) **How many moles** are in 22 grams of argon?

4) **How many grams** are in 88.1 moles of magnesium?

5) **How many moles** are in 2.3 grams of phosphorus?

Mole Conversion Worksheet 1 Answers

A Gutmann

Mole Conversion Worksheet 1 Answers:

Unveiling the Magic of Words: A Report on "**Mole Conversion Worksheet 1 Answers**"

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