

# Modern Physics - Chapter 1

## Chapter 1: Quantum Mechanics

### 1.1 Wave-Particle Duality

- In classical mechanics, particles are thought of as point-like objects with definite positions and velocities.
- However, in quantum mechanics, particles are described by wave functions, which give the probability of finding a particle at a particular position.
- This concept is known as wave-particle duality, which means that particles can behave like waves and vice versa.
- This idea was first introduced by Louis de Broglie in 1924, who proposed that all matter has both particle and wave-like properties.
- The wave-like behavior of particles is often observed in experiments involving diffraction, interference, and tunneling.
- The particle-like behavior of waves is observed in experiments involving the photoelectric effect, where light behaves as if it were composed of particles, or photons.
- The behavior of particles and waves is described by wave functions, which give the probability of finding a particle at a certain position.
- The wave function is a mathematical function that contains all the information about the particle's quantum state, such as its position, momentum, and energy.
- The wave function is often denoted by the symbol  $\psi$  ( $\Psi$ ).
- The wave function is normalized, which means that the total probability of finding the particle at all possible positions is equal to 1.
- The uncertainty principle, which was introduced by Werner Heisenberg in 1927, states that it is impossible to measure both the position and momentum of a particle with absolute precision.
- This is because the act of measuring the position of a particle changes its momentum and vice versa.
- The uncertainty principle is a fundamental principle of quantum mechanics and has important implications for the behavior of particles at the subatomic level.
- In summary, the concept of wave-particle duality explains that particles can behave like waves and waves can behave like particles, and the behavior of particles and waves is described by wave functions, which give the probability of finding a particle at a certain position, and the uncertainty principle states that it is impossible to measure both the position and momentum of a particle with absolute precision.

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