

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds.

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

61.5 yr

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

3 half lives

Average atomic mass


The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ \hline & + & \\ & 63.62 & \rightarrow 64 \end{array}$$

Modern Chemistry Homework Chapteranswers

**John Kenkel, Paul B. Kelter, David S.
Hage**



Modern Chemistry Homework Chapter answers:

Fundamentals of Chemistry: A Modern Introduction Frank Brescia, 2012-12-02 Fundamentals of Chemistry Third Edition introduces the reader to the fundamentals of chemistry including the properties of gases atomic and molecular weights and the first and second laws of thermodynamics Chemical equations and chemical arithmetic are also discussed along with the structure of atoms chemical periodicity types of chemical bonds and condensed states of matter This book is comprised of 26 chapters and begins with a historical overview of chemistry and some terms which are part of the language of chemists Separation and purification are covered in the first chapter while the following chapters focus on atomic and molecular weights stoichiometry the structure of atoms and types of chemical bonds The molecular orbital MO theory of bonding galvanic cells and chemical thermodynamics are considered next Separate chapters are devoted to MO theory of covalent and metallic bonding orbital hybridization intermolecular forces acids and bases ionic equilibrium calculations and polymers and biochemicals This monograph is intended for chemistry students **Chemistry** John Kenkel, Paul B. Kelter, David S. Hage, 2000-09-21 What a great idea an introductory chemistry text that connects students to the workplace of practicing chemists and chemical technicians Tying chemistry fundamentals to the reality of industrial life Chemistry An Industry Based Introduction with CD ROM covers all the basic principles of chemistry including formulas and names chemical bond **Fundamentals of Chemistry: A Modern Introduction (1966)** Frank Brescia, 2012-12-02 Fundamentals of Chemistry A Modern Introduction focuses on the formulas processes and methodologies used in the study of chemistry The book first looks at general and historical remarks definitions of chemical terms and the classification of matter and states of aggregation The text then discusses gases Ideal gases pressure of a gas confined by a liquid Avogadro's Law and Graham's Law are described The book also discusses aggregated states of matter atoms and molecules chemical equations and arithmetic thermochemistry and chemical periodicity The text also highlights the electronic structures of atoms Quantization of electricity spectra of elements quantization of the energy of an electron associated with nucleus the Rutherford Bohr nuclear theory hydrogen atom and representation of the shapes of atomic orbitals are explained The text also highlights the types of chemical bonds hydrocarbons and their derivatives intermolecular forces solutions and chemical equilibrium The book focuses as well on ionic solutions galvanic cells and acids and bases It also discusses the structure and basicity of hydrides and oxides The reactivity of hydrides charge of dispersal and basicity effect of anionic charge inductive effect and basicity and preparation of acids are described The book is a good source of information for readers wanting to study chemistry Essentials of Modern Materials Science and Engineering James A. Newell, 2008-12-15 This book begins with four fundamental tenants The properties of a material are determined by its structure Processing can alter that structure in specific and predictable ways The behavior of materials is grounded in science and is understandable The properties of all materials change over time with use and exposure to environmental conditions When selecting a material sufficient and

appropriate testing must be performed to insure that the material will remain suitable throughout the reasonable life of the product This text assumes that the students are at least sophomores so that they are familiar with basic chemical bonding and the periodic table But it is an introductory materials course so there will be no differential equations percolation theory quantum mechanics statistical thermodynamics or other advanced topics The book is designed as an introduction to the field not a comprehensive guide to all materials science knowledge Instead of going into great detail in many areas the book provides key concepts and fundamentals students need to understand materials science and make informed decisions An example of the philosophy is found in the materials testing section Although countless variations exist in testing techniques the chapter focuses on operating principles and the property to be measured rather than confusing the student with exposition on variations and exceptions That material is beyond the scope of most introductory courses Molecular Mechanics Across Chemistry Anthony K. Rappé, Carla J. Casewit, 1997-05-07 The remarkable breadth of modern molecular mechanics is covered in this textbook developed for an undergraduate or first time course on molecular mechanics The book uses a case study approach designed to give readers exposure to the relevance and utility of molecular mechanics as well as the opportunity to study a particular problem and its solution in depth The remarkable breadth of modern molecular mechanics is covered in this textbook developed for an undergraduate or first time course on molecular mechanics The book uses a case study approach designed to give readers exposure to the relevance and utility of molecular mechanics as well as the opportunity to study a particular problem and its solution in depth **Chemistry in the Modern World** Frank L. Wiseman, 1985 Books in Print Supplement ,2002 Educational Times ,1910 Star Apocrypha Christopher Buckley, 2001 A brilliant expression of philosophy and feeling Christopher Buckley s latest poetry collection explores growing up in the America of the fifties and sixties and coming to terms with the aging process In dazzling language Star Apocrypha bears the beauty and weight of the big questions The poet looks to nature the interior and exterior landscapes for his answers and with wit and high pitched intelligence accepts his art and life in the twenty first century Encyclopaedia Of Modern Tech. K.P. Yadav, Malti Sundram, Malti Sundram, 2004 Education Outlook ,1910 **Children's Books in Print** R R Bowker Publishing, Bowker, 1999-12 **Educational Times and Journal of the College of Preceptors** ,1910

Polymer Chemistry Timothy P. Lodge, Paul C. Hiemenz, 2020-07-14 A well rounded and articulate examination of polymer properties at the molecular level Polymer Chemistry focuses on fundamental principles based on underlying chemical structures polymer synthesis characterization and properties It emphasizes the logical progression of concepts and provide mathematical tools as needed as well as fully derived problems for advanced calculations The much anticipated Third Edition expands and reorganizes material to better develop polymer chemistry concepts and update the remaining chapters New examples and problems are also featured throughout This revised edition Integrates concepts from physics biology materials science chemical engineering and statistics as needed Contains mathematical tools and step by step derivations for

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