



Manual Galloper Super Exceed

Liying Dong



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web e angular acceleration or deceleration of the motor in rolling s 40 60 rpm sec j moment of inertia of rotating parts referred to motor shaft g mass of rotating parts unit d diameter of gyration g acceleration due to gravity i e 9 81 m sec² s angular acceleration for electric motor of rolling mill the value varies from

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web 1 consider the reaction $\text{i}_2\text{o}_5 + 5 \text{co} \rightarrow 5 \text{co}_2 + \text{i}_2$ g a 80 0 grams of iodine v oxide i_2o_5 reacts with 28 0 grams of carbon monoxide co determine the mass of iodine i_2 which could be produced b if in the above situation only 0 160 moles of iodine i_2 was produced i what mass of iodine was produced

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$$2\text{Cu} + \text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu}(\text{NO}_3)_2$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$
 $8.3 \text{ mole } \text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web percent yield is a measure of the efficiency of the experimental design yield efficiency $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
 4 Al + 3 $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
 b determine the theoretical yield of Al_2O_3 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} = 5.93 \text{ mol Al}$
 $\frac{5.93 \text{ mol Al}}{4} \times 2 = 2.96 \text{ mol } \text{Al}_2\text{O}_3$
 $2.96 \text{ mol } \text{Al}_2\text{O}_3 \times 101.96 \text{ g/mol} = 302.3 \text{ g } \text{Al}_2\text{O}_3$
 c determine the percent yield $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7 Na + 2 C_2O_4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4.33 g of H_2 are left over 13

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web the percent yield is calculated as follows text percent yield $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$ percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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