

Assessment

Chapter Test B

Atomic Physics

MULTIPLE CHOICE

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- What is the frequency of a photon with an energy of $1.99 \times 10^{-19} \text{ J}$?
($h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$)
 - $1.00 \times 10^{14} \text{ Hz}$
 - $2.00 \times 10^{14} \text{ Hz}$
 - $3.00 \times 10^{14} \text{ Hz}$
 - $4.00 \times 10^{14} \text{ Hz}$
- Light with an energy equal to three times the work function of a given metal causes the metal to eject photoelectrons. What is the ratio of the maximum photoelectron kinetic energy to the work function?
 - 1 : 1
 - 2 : 1
 - 3 : 1
 - 4 : 1
- A monochromatic light beam with a quantum energy value of 3.0 eV is incident upon a photocell. The work function of the photocell is 1.6 eV. What is the maximum kinetic energy of the ejected electrons?
 - 4.6 eV
 - 4.8 eV
 - 1.4 eV
 - 2.4 eV
- What causes the bright lines in the emission spectrum of an element to occur?
 - Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - Photons are emitted when electrons jump from a lower-energy to a higher-energy state.
- What causes the dark lines in the absorption spectrum of an element to occur?
 - Photons are absorbed when electrons jump from a higher-energy to a lower-energy state.
 - Photons are emitted when electrons jump from a higher-energy to a lower-energy state.
 - Photons are absorbed when electrons jump from a lower-energy to a higher-energy state.
 - Photons are emitted when electrons jump from a lower-energy to a higher-energy state.

Holt Physics Chapter 17 Test Answers

A Loxley



Holt Physics Chapter 17 Test Answers:

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