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**Chapter 8 Atomic Structure Section 8.1 Studying Atomic (pages 189-195)** This section explains how atomic models evolved. **Reading Strategy (page 190)** Summarizing As you read, fill in the table about atomic models. For more details on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook. **Atomic Models Scientific Evidence Model** Model of atoms in compounds Deducted from Rutherford's Positive, dense nucleus. **Atomic Greek Models of Atoms (page 191)** 1. Democritus called the smallest particles of matter because they could not be split. 2. He saw the four elements that Aristotle used in his model of matter: a, b, c, d. Dalton's Atomic Theory (page 193) 3. Thomson's Model

Model	Structure
Democritus's Model	Democritus's model of matter was based on the idea that matter is made of tiny particles called atoms. Atoms cannot be created or destroyed. Atoms are indivisible particles, which cannot be created or destroyed.
Aristotle's Model	Aristotle's model of matter was based on the idea that matter is made of four elements: earth, air, fire, and water. He thought that these elements were made of tiny particles called atoms.
Thomson's Model	Thomson's model of matter was based on the idea that matter is made of tiny particles called atoms. Atoms are made of protons and electrons. Protons are positively charged and electrons are negatively charged. The protons and electrons are held together by the force of attraction.
Rutherford's Model	Rutherford's model of matter was based on the idea that matter is made of tiny particles called atoms. Atoms are made of protons and neutrons. Protons are positively charged and neutrons are neutral. The protons and neutrons are held together by the force of attraction.
Bohr's Model	Bohr's model of matter was based on the idea that matter is made of tiny particles called atoms. Atoms are made of protons and electrons. Protons are positively charged and electrons are negatively charged. The protons and electrons are held together by the force of attraction. Electrons move in fixed orbits around the nucleus.

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For more details on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook.

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**Yi-Tong Ma**



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