



COVALENT BONDING

Practice Problems

Do your best to do better than your previous problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following compounds as ionic or covalent.

- a. Na_2SO_4 b. CH_3NH_2 c. HCl d. H_2O
 e. KNO_3 f. H_2SO_4 g. H_2O_2 h. H_2O

2. Write the formula for the following molecular compounds.

- a. N_2O_5 b. N_2O_4 c. N_2O_3
 d. N_2O e. N_2O_2 f. N_2O_6

3. Write the Lewis structure for each of the following molecular compounds.

4. Write the Lewis structure for each of the following molecular compounds. Be sure to show all unshared valence electrons and indicate the formal charge on each atom.

5. Write the Lewis structure for each of the following molecular compounds. Be sure to show all unshared valence electrons and indicate the formal charge on each atom.

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF .



2. Draw the Lewis structure for hydrogen peroxide, H_2O_2 .



3. Draw the Lewis structure for carbon dioxide, CO_2 .



4. Draw the Lewis structure for acetylene, C_2H_2 .

5. Draw the Lewis structure for each of the following molecular compounds. Be sure to show all unshared valence electrons and indicate the formal charge on each atom.



6. Draw the Lewis structure for each of the following molecular compounds. Be sure to show all unshared valence electrons and indicate the formal charge on each atom.

SECTION 8.3 BONDING: THE ORBITAL

1. Write the Lewis structure for each of the following molecular compounds. Be sure to show all unshared valence electrons and indicate the formal charge on each atom.

2. Write the Lewis structure for each of the following molecular compounds. Be sure to show all unshared valence electrons and indicate the formal charge on each atom.

Guided Practice Problem Chapter 8 Covalent Bonding

Arshad Iqbal



Guided Practice Problem Chapter 8 Covalent Bonding:

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