

Chapter 6: Chemical Bonding Guided Notes

I. Why do Atoms Combine?

A. Chemical Formula

1. The chemical formula shows:
 - a. The _____ in the compound.
 - b. The ratio of their _____

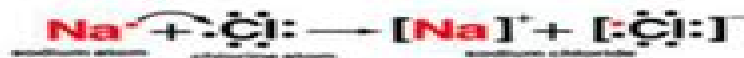


B. Chemical Bond

1. Strong attractive force between _____ or ions in a molecule or compound.
2. Formed by:
 - a. _____ electrons, (e-) by losing or gaining electrons.
 - b. _____ electrons (e-)

C. Stability

1. Octet Rule
 - a. Most atoms form bonds in order to have _____ valence electrons (e-).
 - b. _____ outer energy level.
 - c. Want to be like the Noble Gases (group 18).
2. Stability is the driving force behind bond formation. All atoms bond to become stable.
3. Transferring e-



4. Sharing e-



II. Kinds of Chemical Bonds

A. Ionic Bond

1. Attraction between 2 _____ charged ions.
2. Ions are _____ atoms.
3. A _____ is a _____ charged ion. Forms when an atom loses an electron.
Formed by _____.
4. A _____ is a _____ charged ion. Forms when an atom gains an electron.
Formed by _____.
5. Formed by _____ electrons (e-) from a _____ to a _____.
6. Ions form a 3-D crystal lattice.



Guided Notes Ionic Bonds Chapter 6

Sebastian Brünink

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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