

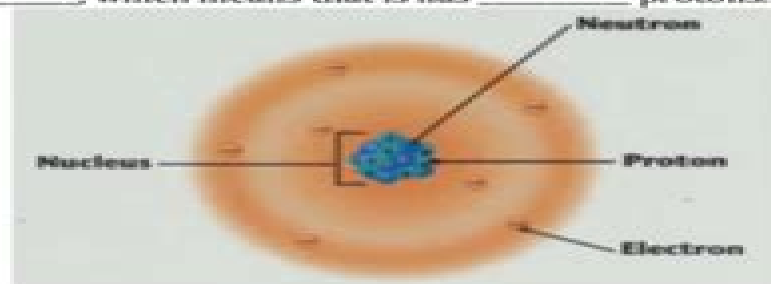
## Physical Science Standards:

**SPS1:** Students will investigate our current understanding of the atom.

### a. Examine the structure of the atom in terms of

- proton, electron, and neutron locations.
  - Proton has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Protons are also the elements' \_\_\_\_\_ number of the identity of the element, so protons are found by looking at the atomic number
  - Neutron has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Neutrons are found by subtracting the \_\_\_\_\_ Number from the \_\_\_\_\_ Mass or the BIG #- SMALL #.
  - Electron has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Electrons are equal to the number of \_\_\_\_\_ in a neutral atom.
  - The nucleus is made up of these 2 particles, \_\_\_\_\_ and \_\_\_\_\_.
  - The electrons are located in this region of the atom, known as the \_\_\_\_\_.
  - The electron cloud is divided into energy levels or energy shells, the first level will hold \_\_\_\_\_ electrons and the second level will hold \_\_\_\_\_ electrons.
  - The valence shell or valence level is the \_\_\_\_\_ energy shell of the atom.
- atomic mass and atomic number.
  - Atomic number is equal to the number of \_\_\_\_\_ in the atom and is also the atom's identity—it's the small number of the periodic table.
  - Atomic mass is equal to the average mass of all the isotopes of a particular element—it's the big number on the periodic table.
- atoms with different numbers of neutrons (isotopes).
  - Define Isotope \_\_\_\_\_.
  - Ex. Carbon-12 and Carbon-14
    - Carbon 12 has \_\_\_\_\_ P \_\_\_\_\_ N \_\_\_\_\_ E
    - Carbon 14 has \_\_\_\_\_ P \_\_\_\_\_ N \_\_\_\_\_ E
- explain the relationship of the proton number to the element's identity.
  - The \_\_\_\_\_ number is equal to the number of \_\_\_\_\_ in the atom which is how the element is identified.
  - Ex. Nitrogen has an atom number of \_\_\_\_\_, which means that it has \_\_\_\_\_ protons.

molybdenum	← element name
42	← atomic number number of protons (Z)
Mo	← atomic symbol
95.94	← atomic mass A (this is an average mass)



- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number - Atomic Number

# Eoct Physics Study Guide

**A Gutmann**



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## **Table of Contents Eoct Physics Study Guide**

1. Understanding the eBook Eoct Physics Study Guide
  - The Rise of Digital Reading Eoct Physics Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Eoct Physics Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Eoct Physics Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Eoct Physics Study Guide
  - Personalized Recommendations
  - Eoct Physics Study Guide User Reviews and Ratings
  - Eoct Physics Study Guide and Bestseller Lists
5. Accessing Eoct Physics Study Guide Free and Paid eBooks
  - Eoct Physics Study Guide Public Domain eBooks
  - Eoct Physics Study Guide eBook Subscription Services

- Eoct Physics Study Guide Budget-Friendly Options
- 6. Navigating Eoct Physics Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Eoct Physics Study Guide Compatibility with Devices
  - Eoct Physics Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Eoct Physics Study Guide
  - Highlighting and Note-Taking Eoct Physics Study Guide
  - Interactive Elements Eoct Physics Study Guide
- 8. Staying Engaged with Eoct Physics Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Eoct Physics Study Guide
- 9. Balancing eBooks and Physical Books Eoct Physics Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Eoct Physics Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Eoct Physics Study Guide
  - Setting Reading Goals Eoct Physics Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Eoct Physics Study Guide
  - Fact-Checking eBook Content of Eoct Physics Study Guide
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- 
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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