

2.1.4 & 2.1.5 - Acids and redox

What are standard solutions?	A solution for which its concentration is known accurately.
What 3 things should you ensure when making standard solutions?	<ol style="list-style-type: none"> 1. Ensure you have clean dry equipment. 2. Make washings to remove any salt BEFORE MAKING up the solution to the specified amount. 3. Invert the volumetric flask to ensure everything is dissolved.
How does ammonia react with acid?	<ul style="list-style-type: none"> • ammonia + acid \rightarrow ammonium salt • $NH_3(aq) + HNO_3(aq) \rightarrow NH_4NO_3(aq)$
What are strong and weak acids ? Give 3 strong acids and 1 weak acid	<ul style="list-style-type: none"> • Strong acids dissociate fully to release H^+ ions in a solution e.g. $HCl \rightarrow H^+ + Cl^-$ • Weak acids dissociate partially to release H^+ ions in a solution – exist in equilibrium • Strong acids include: HCl, H_2SO_4 and HNO_3. • Weak acids include carboxylic acids such as CH_3COOH.
What are strong and weak bases? Give 1 weak base	<ul style="list-style-type: none"> • Strong bases dissociate fully to release OH^- ions in a solution • Weak bases dissociate partially to release OH^- ions in a solution – exist in equilibrium e.g. NH_3
What are diprotic and triprotic acids (with examples) and what does this mean for neutralisation?	<ul style="list-style-type: none"> • Diprotic acids and triprotic acids donate two and three protons respectively. • Examples include: H_2SO_4 and H_3PO_4. • They need to react to the number of moles of base to be neutralised

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Table of Contents Mark Scheme For May 2013 Chemistry Ocr

1. Understanding the eBook Mark Scheme For May 2013 Chemistry Ocr
 - The Rise of Digital Reading Mark Scheme For May 2013 Chemistry Ocr
 - Advantages of eBooks Over Traditional Books
2. Identifying Mark Scheme For May 2013 Chemistry Ocr
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mark Scheme For May 2013 Chemistry Ocr
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mark Scheme For May 2013 Chemistry Ocr
 - Personalized Recommendations
 - Mark Scheme For May 2013 Chemistry Ocr User Reviews and Ratings

- Mark Scheme For May 2013 Chemistry Ocr and Bestseller Lists
- 5. Accessing Mark Scheme For May 2013 Chemistry Ocr Free and Paid eBooks
 - Mark Scheme For May 2013 Chemistry Ocr Public Domain eBooks
 - Mark Scheme For May 2013 Chemistry Ocr eBook Subscription Services
 - Mark Scheme For May 2013 Chemistry Ocr Budget-Friendly Options
- 6. Navigating Mark Scheme For May 2013 Chemistry Ocr eBook Formats
 - ePub, PDF, MOBI, and More
 - Mark Scheme For May 2013 Chemistry Ocr Compatibility with Devices
 - Mark Scheme For May 2013 Chemistry Ocr Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mark Scheme For May 2013 Chemistry Ocr
 - Highlighting and Note-Taking Mark Scheme For May 2013 Chemistry Ocr
 - Interactive Elements Mark Scheme For May 2013 Chemistry Ocr
- 8. Staying Engaged with Mark Scheme For May 2013 Chemistry Ocr
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mark Scheme For May 2013 Chemistry Ocr
- 9. Balancing eBooks and Physical Books Mark Scheme For May 2013 Chemistry Ocr
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mark Scheme For May 2013 Chemistry Ocr
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mark Scheme For May 2013 Chemistry Ocr
 - Setting Reading Goals Mark Scheme For May 2013 Chemistry Ocr
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mark Scheme For May 2013 Chemistry Ocr
 - Fact-Checking eBook Content of Mark Scheme For May 2013 Chemistry Ocr
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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