

23. Limiting Reactant and Excess Reactant

Background

Limiting Reactant and Excess Reactant

In reactions where the reactants are not present in a stoichiometric ratio, the amount of product produced is determined by the amount of the reactant that is fully consumed in the reaction first. This reactant is called the limiting reactant (or limiting reagent). It is the amount of this reactant that is used to calculate the amount of product, just like a baked recipe can't be adjusted based on the amount of an ingredient that's in short supply.

Mass and Molar Ratios in Chemical Reactions

Reactions occur according to the number of atoms and molecules that are available to react in a precise ratio. The law of conservation of mass states that matter can be neither created nor destroyed. Therefore, the number of atoms of each element on the reactants side must be the same as the number of atoms of each element on the products side. This quantitative study of chemical reactions is called stoichiometry. The subscripts in a balanced reaction are related to each other by stoichiometric molar ratios. For example, nitrogen reacts with hydrogen to form ammonia in a 1:3 stoichiometric ratio in the balanced equation:



The molar ratios can be used in conversion factors to find the amount in moles of one substance given the amount of another.

Laboratory Manual Limiting Reactant

Thomas Griffiths



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