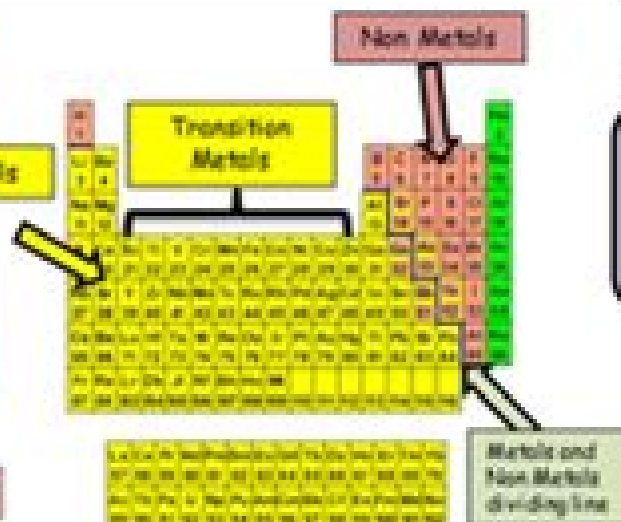


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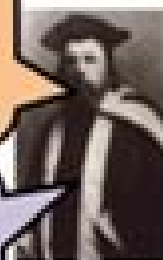


Particle	Relative Mass	Relative Charge
Proton	1	+
Neutron	1	0
Electron	Very small	-



I came up with the first periodic table. I arranged the elements in order of increasing atomic mass

Dmitri Mendeleev



I put elements with similar properties into horizontal rows. I also left gaps incase any new elements were discovered

He

4 ← **MASS NUMBER** = number of protons + number of neutrons

2 ← **PROTON NUMBER** = number of protons (atomic)

CHARGE

Isotopes

Higher

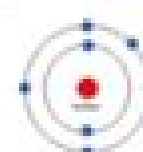
Arrange in a table with a different number of neutrons

Notice that the mass number is different. How many neutrons does each isotope have?

16 17 18

Each isotope has 8 protons - if it didn't then it just wouldn't be oxygen any more

14
N
7



Electron structure
+ 2, 5

24
Mg
12

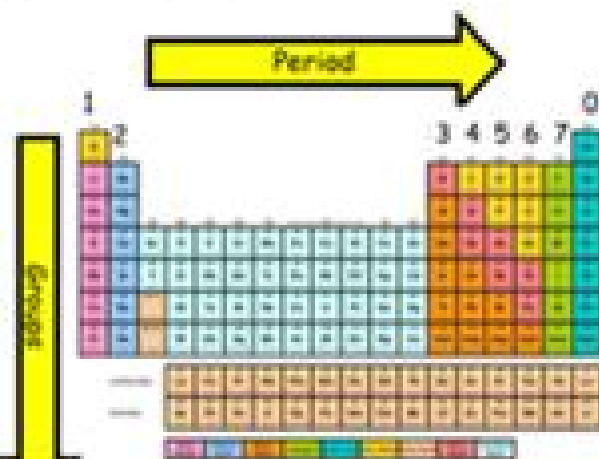


Electron structure
+ 2, 8, 2

40
Ca
20



Electron structure
+ 2, 8, 8, 2



How can an atom have a decimal for its mass?

This is because out of every four naturally occurring chlorine atoms, 3 have a mass of 35 and 1 has a mass of 37 so the average atomic mass is:

$$35 \times 3 + 37 \times 1 = 35.5$$

Elements in the same group have the same number of electrons in the outer shell (this corresponds to their group number)



(Most important) All of the elements in the same group have similar **PROPERTIES**. This is how I thought of the periodic table in the first place. This is called **PERIODICITY**.

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L. D. R. WILFORD



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web anatole france kitap açıklaması kudretli eşitliğiyle yasa zenginlerin de fakirlerin de köprülerin altında uyumasını sokaklarda dilenmesini ve ekmek çalmasını yasaklar anatole france bu adam anatole france tır Üslubunda

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web anatole france 16 nisan 1844 tarihinde fransa da dünyaya geldi bir kitapçının oğlu olduğu için daha küçük yaşlarında kitaplarla tanıştı okulunda da hümanist kültürle yetiştikten sonra edebiyatla ilgilenmeye karar verdi

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