F321

Question		Answer	Marks	Guidance
2 (a)	(1)	AJ ³¹ ✓ SO ₄ ³⁻ ✓	2	2
	(11)	Al ₂ O ₃ (s) + 3H ₂ SO ₄ (aq) → Al ₂ (SO ₄) ₃ (aq) + 3H ₂ O(I) Correct species AND correctly balanced ✓ state symbols on correct species ✓	2	ALLOW multiples
	(m)	(The number of) water(s) of crystallisation ✓	10	IGNORE hydrated OR hydrous OR 'contains water'
	(hr)	First check the answer on the answer line. If answer = 16, award 3 marks Correctly calculates amount of Al ₂ (SO ₄) ₃ : 6.846 / 342.3 = 0.02(00) mol Correctly calculates amount of H ₂ O: 5.760 / 18.0 = 0.32(0) mol Correctly calculates whole number ratio of mol of H ₂ O: Al ₂ (SO ₄) ₃ to give x = 16 Y	3	If there is an alternative answer, check to see if there is any ECF credit possible using working below ALLOW as ECF from 12:606/342.3 = 0.0368(273) AND 0.32/0.0368(273) To give x = 9 for two marks ALLOW calculator value or rounding to 2 significant figures or more BUT IGNORE 'traiting' zeroes, eg 0.200 allowed as 0.2. ALLOW ECF for calculation of correctly rounded whole number value of H ₂ O from incorrect mol of H ₂ O and / or incorrect mol of Al ₂ (SO ₄) ₃ BUT x must be a whole number ALLOW alternative method Mol of Al ₂ (SO ₄) ₃ : 6.846 / 342.3 = 0.02(00) mol (first mark) Molar mass of Al ₂ (SO ₄) ₃ : xH ₂ O: 12:606 / 0.02(00) = 630.3 g mol ⁻¹ (second mark) Mass of water per mol = 630.3 – 342.3 = 288 AND 288/18 to give x = 16 (third mark)

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