

LONDON COLLEGE OF INTERNATIONAL CHEMISTS – 6000 (004)

NAME: _____ COURSE: _____

ENTRY NUMBER: _____ ALPH. NO. _____

ICM: IDEAS (CONCEPTS IN INORGANIC CHEMISTRY)

ICM 11

Class : ICM-004-004
 Date : 2022 - 2-2022

Page: No. _____

Time : 1:00 Minutes

PART – A

20 x 2 = 40 marks

Answer ALL the questions.

1. State three bonding principles.
2. Describe the structure of CH_4 in terms of the VSEPR theory.
3. Write the chemical formula for CO_2 in the VSEPR.
4. Describe the structure of CO_2 using VSEPR theory.
5. What is the relationship between the number of valence electrons of an element and the number of valence electrons of an element in forming the molecule of an ionic compound? Explain.
6. Describe and explain how the bond angles in the molecule H_2O are affected by the lone pairs.
7. What are the three types of hybridization?
8. What are the three types of hybridization?
9. Describe the structure of CO_2 in the VSEPR.
10. What are the three types of hybridization?

PART – B

20 x 2 = 40 marks

Answer any FIVE the questions.

11. What are the three types of hybridization? Explain the structure of CO_2 in the VSEPR.
12. Describe the structure of CO_2 in the VSEPR.
13. Describe the structure of CO_2 in the VSEPR.
14. Describe the structure of CO_2 in the VSEPR.
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21. Describe the structure of CO_2 in the VSEPR.
22. Describe the structure of CO_2 in the VSEPR.

PART – C

20 x 2 = 40 marks

Answer any FIVE the questions.

23. What are the three types of hybridization? Explain the structure of CO_2 in the VSEPR.
24. Describe the structure of CO_2 in the VSEPR.
25. Describe the structure of CO_2 in the VSEPR.
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