

Lab 10 – pH and Buffers

A. Objectives:

1. Constructive proper titration and titration curve.
2. Understand and explain how the pH scale works.
3. Understand and describe the concept of buffer capacity and strength.
4. Compare acids and bases in terms of their strength.
5. Calculate the pH of various buffer solutions.
6. Calculate equilibrium pH, buffer capacity and buffer strength of a buffer.

B. Background: Titration, Equilibrium and Coupling

1. Titration:

Chemical titration is a quantitative method for determining the concentration of a solution containing a known volume of a solution of a known concentration. It is a technique for determining the concentration of a solution of a known volume of a known concentration.

2. Buffering and titration: A buffer is a solution of a weak acid and its conjugate base. The buffer capacity is the ability of a buffer to resist changes in pH. The buffer capacity is the ability of a buffer to resist changes in pH. The buffer capacity is the ability of a buffer to resist changes in pH. The buffer capacity is the ability of a buffer to resist changes in pH.

3. Buffering and titration: A buffer is a solution of a weak acid and its conjugate base. The buffer capacity is the ability of a buffer to resist changes in pH. The buffer capacity is the ability of a buffer to resist changes in pH. The buffer capacity is the ability of a buffer to resist changes in pH. The buffer capacity is the ability of a buffer to resist changes in pH.

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Concentration (molar)	pH	Concentration	pH	Concentration	pH
0.100 M		0.100 M		0.100 M	
0.050 M		0.050 M		0.050 M	
0.025 M		0.025 M		0.025 M	

Lab Manual Bio 150 Dccc

Siddappa N.Byrareddy



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