

Name: \_\_\_\_\_

**CHEMISTRY 101**  
**EXAM 3 FORM B****SECTIONS 501-511****SPRING 2005**  
**DR. KEENEY-KENNICUTT**

- Directions:
- (1) Put your name and signature on PART 2 of the exam where indicated.
  - (2) Sign the Aggie Code on PART 2 of this exam.
  - (3) Each multiple choice question is actually 2 questions on your scanning sheet. If you are sure of an answer, put the same answer down for both questions for 5 pts. If you cannot decide between two answers, put one answer down for one question and the other answer down for the other question. If you get one correct you'll get half credit for 2.5 pts. If there is an ambiguous multiple choice question, use the last page to explain your answer.
  - (4) Do NOT write on the envelope.
  - (5) When finished, put everything in the envelope and wait to be excused. At the table, take everything out of the envelope. You can pick up the multiple choice part with the answers outside my office after the exam.
  - (6) There are a total of 27 questions (15 actual questions).

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**PART 1**

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**1&2.** The number of protons and electrons found in the aluminum ion that is isoelectronic with a noble gas is:

- (a) 13 p and 16 e<sup>-</sup>                      (b) 10 p and 13 e<sup>-</sup>                      (c) 13 p and 10 e<sup>-</sup>  
(d) 13 p and 13 e<sup>-</sup>                      (e) some other combination

**3&4.** Which of the following bonds can be classified as polar covalent?

- (a) Cu-Ti                      (b) H-F                      (c) S-I                      (d) C-S                      (e) Na-F

**5&6.** Which indication of relative acid strength is WRONG?

- (a)  $\text{H}_2\text{S} < \text{H}_2\text{Se}$                       (b)  $\text{H}_2\text{SeO}_4 < \text{H}_2\text{SO}_4$                       (c)  $\text{HBrO}_2 > \text{HBrO}_3$   
(d)  $\text{H}_2\text{SO}_4 > \text{HSO}_4^-$                       (e)  $\text{HF} < \text{HI}$

# Guide For Chemistry Spring Exam

**WJ Hussar**



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