

GCSE Mathematics Practice Tests: Set 2

Paper 2H (Calculator)

Time: 1 hour 30 minutes

You should have: Ruler graduated in centimetres and millimetres, protractor, pair of compasses, pen, HB pencil, eraser, calculator.

Instructions

- Use **black** ink or ball-point pen.
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions.
- Answer the questions in the spaces provided – *there may be more space than you need.*
- **Calculators may be used.**
- Diagrams are **NOT** accurately drawn, unless otherwise indicated.
- You must **show all your working out.**



Information

- The total mark for this paper is 80.
- The marks for **each** question are shown in brackets – *use this as a guide as to how much time to spend on each question.*

Advice

- Read each question carefully before you start to answer it.
- Keep an eye on the time.
- Try to answer every question.
- Check your answers if you have time at the end.

Gcse Edexcel Maths Mock Paper 2014 H1

Michael Seilmaier

A red circular graphic with a gradient, appearing as a semi-circle or a stylized arrow pointing to the right, located to the right of the author's name.

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web limiting reactant theoretical and percentage yields key $4 \text{ko}_2 + 2 \text{h}_2\text{o} \rightarrow 4 \text{koh} + \text{o}_2$ g if a reaction vessel contains 0 mol ko_2 and 0 mol h_2o what is the limiting reactant

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$$2\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{CuNO}_3 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} = 8.3 \text{ mole } \text{H}_2\text{O}$
 $8.3 \text{ mole } \text{H}_2\text{S} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$
 problems 6 & 8 are similar to problem 5

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 $21.3 \text{ g } \text{O}_2 \times \frac{1 \text{ mole } \text{O}_2}{32 \text{ g } \text{O}_2} = 0.666 \text{ moles } \text{O}_2$

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
 4 Al + 3 $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
 b determine the theoretical yield of Al_2O_3
 $160.0 \text{ g Al} \times \frac{1 \text{ mole Al}}{27 \text{ g Al}} \times \frac{2 \text{ mole } \text{Al}_2\text{O}_3}{4 \text{ mole Al}} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 150.1 \text{ g } \text{Al}_2\text{O}_3$
 c determine the percent yield
 $\frac{100.26 \text{ g } \text{Al}_2\text{O}_3}{150.1 \text{ g } \text{Al}_2\text{O}_3} \times 100 = 66.8\%$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of h 2 are left over 13

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