



I. DALTON'S ATOMIC THEORY

- John Dalton proposed an atomic theory of matter that can explain chemical observation as predicted by the fundamental laws:

1. **Law of Conservation** states that matter is not created nor destroyed.
2. **Law of definite proportion** states that different samples of the same compound always contain the constituent elements in the same proportion by mass.
3. **Law of multiple proportion** states that if two elements form more than one compound, the various masses of one element combining with a fixed mass of another element are related by small whole-number ratios.

ATOMIC THEORY

1. Atoms are the smallest particles of matter. They cannot be divided into smaller particles, they also cannot be created nor destroyed.
2. All atoms of an element are identical, but the atoms of one element are different from atoms of other elements in terms of mass, size, and properties.
3. Compounds are composed of atoms of more than one element which are combined in fixed ratios.
4. Atoms retain their identity during chemical reactions, which involve combination, separation, and rearrangement. They are indestructible.

Matter And Atoms Study Guide

EW Minium



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Matter And Atoms Study Guide Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Matter And Atoms Study Guide**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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