INTRODUCTION

The purpose of this experiment is to determine the pH of various concentrations of acids, bases, salts, and buffer solutions. The data will be used to determine the acid-base nature in the hydrolysis of several salts and behavior of buffer solutions. Recall that salts are the product of neutralization reactions between acids and bases, generically shown below with an acid, HA, and base, B.

Eq. 1
$$HA(aq) + BOH(aq) \leftrightarrow BA(aq) + H_1O(l)$$

As nearly all salts are strong electrolytes, they exist in solution as ions. Marry of these ions then can react with water to produce an acidic or basic solution. This reaction of salts with water is called a hydrolysis reaction. An anion, such as the generic A, is the conjugate base of the weak acid HA. These anions (A) will then react with water to generate OH.

Eq. 2 A
$$(Bq)$$
 + H,O(I) \leftrightarrow HA(Bq) + OH (Bq)

Cations such has NH₄* or highly charged metals like Fe³⁺ come from weak bases and react with water to form H* (H₂O*, hydronium) ions.

HYDROLYSIS OF ANIONS: BASIC SOLUTIONS

First, let us focus on the hydrolysis of anions. For example, the nitrite ion (NO₂), which will react with water (like a weak base, for example NH₃) to generate hydroxide and thus a basic solution (pH greater than 7):

Hydrolysis Of Salts Lab Report Answers

Ying-Ying Zheng

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