1. What volume of hydrogen gas at STP is produced from the reaction of 50.0g of Mg and 75.0 grams of HCl? How much of the excess reagent is left over (in grams)?

$$Mg(s) + 2\,HCl(aq) \rightarrow MgCl_2(s) + H_2(g)$$

Do a standard mass to mass problem starting with each reactant

$$50.0\ g\,Mg \times \frac{1\,mol\,Mg}{24.31\,g\,Mg} \times \frac{1\,mol\,H_2}{1\,mol\,Mg} \times \frac{22.4\,L}{1\,mol\,H_2} = 46.1\ L\ H_2$$

$$75.0\ g\,HCl \times \frac{1\,mol\,HCl}{36.46\,g\,HCl} \times \frac{1\ mole\ H_2}{2\,mol\,HCl} \times \frac{22.4\,L}{1\,mol\,H_2} = 23.0\ L\ H_2$$

HCl is the limiting reactant!!

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SA Dillow

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