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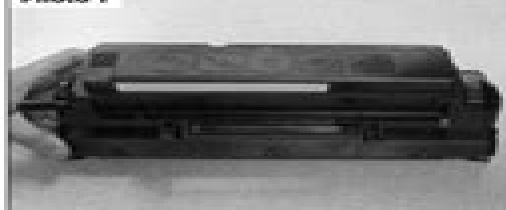
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Needlenose Pliers
Soft Lint-Free Cloth
Cotton Swabs
Anhydrous Isopropyl Alcohol
Vacuum or Compressed Air

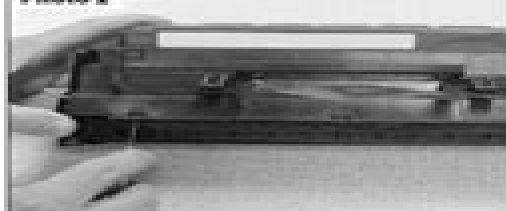
Photo 1



Step 1

Using a Phillips screwdriver remove the two screws from the top of the cartridge.
(See Photo 1)

Photo 2



Step 2

With a small flat blade screwdriver release the three clips located on the front of the cartridge under the handle.
(See Photo 2)

Photo 3



Step 3

To separate the two sections, lift up on the front of the toner hopper. Be careful when separating the two sections, toner will spill out of the toner hopper.
(See Photo 3)

NOTE: Do not break off the clips on the back of the hopper.

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G Psacharopoulos



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web this volume is known as the molar gas volume V_m the volume of a gas may be calculated from its number of moles using volume of gas moles V_m rearranging this moles of gas $\frac{V}{V_m}$

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web jun 8 2023 $V = 1.00 \text{ mol} \times 0.08206 \text{ L atm mol}^{-1} \text{ K}^{-1} \times 273 \text{ K} = 22.4 \text{ L}$ in other words 1 mole of a gas will occupy 22.4 L at STP assuming ideal gas behavior at STP the volume of a gas is only dependent on number of moles of

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web aug 3 2022 volumes of gases and liquids are measured in cubic centimetres cm^3 or cubic decimetres dm^3 or litres L 1 mole of any gas at room temperature and pressure occupies 24 dm^3 $24\,000 \text{ cm}^3$ this value is known as the molar volume

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