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so calculating enthalpy change of reaction in 2 $\Delta H_{\text{reaction}} = 6 \times \Delta H_{\text{f}}^\circ \text{ CH}_4(\text{g}) - 6 \times \Delta H_{\text{f}}^\circ \text{ C}_2\text{H}_2(\text{g})$ $6 \times 523 - 6 \times 82 = 3140 - 492 = 2648 \text{ kJ}$

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