

Name: _____ Date: _____ Lab Section: _____

Experiment 2: Molecular Models REPORT SHEET

Pre-lab Study Questions:

- Objectives of the experiment (in your own words):
- What kinds of electrons form the covalent bonds in molecules?
- Is a polar molecule symmetrical or asymmetrical?

FORMULA OF MOLECULE	LEWIS STRUCTURE	3-D SHAPE w/ lone pairs	MOLECULAR GEOMETRY	MOLECULAR POLARITY
example: CH_4	$\begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{H} \\ \\ \text{H} \end{array}$		tetrahedral	nonpolar
H_2O	$\begin{array}{c} \cdot\cdot \\ \cdot\cdot \\ \text{H} \cdot \text{O} \cdot \text{H} \\ \cdot\cdot \\ \cdot\cdot \end{array}$	$\begin{array}{c} \text{H} \cdot \cdot \cdot \cdot \text{O} \cdot \\ \cdot\cdot \\ \text{H} \end{array}$	tetrahedral	polar
H_2O^+	$\text{H} : \ddot{\text{O}} : \text{H}$	$\begin{array}{c} \text{H} \cdot \cdot \cdot \cdot \text{O} \cdot \cdot \cdot \cdot \text{H} \\ \cdot\cdot \\ \cdot\cdot \end{array}$	tetrahedral	nonpolar
C_2H_6				
C_2H_4	$\begin{array}{c} \text{H} & & \text{H} \\ & \backslash & / \\ & \text{C} = \text{C} \\ & / & \backslash \\ \text{H} & & \text{H} \end{array}$		trigonal planar	non polar
C_2H_2				
CO_2	$\ddot{\text{O}} = \text{C} = \ddot{\text{O}}$		Linear	Non polar
H_2O_2				
CHCl_3				

Geometrical Structure Of Molecules Lab Answers

Massachusetts Agricultural College



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