LOHS Honors Chemistry Balancing Chemical Equations

Part I: Balance these chemical equations. Record what type of reaction on the blank in the right margin.

 BaCl₂ + AKClO₃ → Ba(ClO₃)₂ + AKCl 	_DR
2. $Pb(OH)_2 \rightarrow PbO + H_2O$	D
3. $2Na + 2H_2O \rightarrow 2NaOH + H_2$	SR
4. $Pb(NO_3)_2 + K_2CrO_4 \rightarrow PbCrO_4 + 2KNO_3$	DR
5. $Al_2(SO_4)_3 + 3Ca(OH)_2 \rightarrow 2Al(OH)_3 + 3CaSO_4$	DR
6. $Mg + 2HCl \rightarrow MgCl_2 + H_2$	SR
7. $Sn + 2KOH \rightarrow K_2SnO_2 + H_2$	SR
8. 2 PaI ₅ → 2Pa + 5I ₂	
9. $3SiF_4 + 3H_2O \rightarrow 2H_2SiF_6 + H_2SiO_3$	SR
10. $Cl_2O_7 + H_2O \rightarrow AHClO_4$	
11. $/6$ Na + S ₈ \rightarrow 8 Na ₂ S	S
12. N ₂ O ₃ + H ₂ O → 2HNO ₂	s

Part II: Write balanced chemical equations for each of the reactions below. Assume that all ionic compounds are dissolved in water (aqueous). Also be sure include physical states where appropriate.

1. Iron metal combines with oxygen gas to form iron (III) oxide solid

Chromium metal and sulfuric acid (H2SO4) react to form chromium (III) sulfate solution and hydrogen gas.

 Barium chromate solution and aqueous hydrochloric acid (HCl) react to produce barium chloride solution and hydrogen chromate (solid).

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Tara C. Dale, Mandi S. White

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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 Let's Review Regents:

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