

Procedure for making NaCl-Aid (for every lab group of 3-4 people):

1. Using your calculations from question #4 above, measure out the correct amount of solid NaCl-Aid powder (grams) into the assigned cup to make a 0.1 M solution. **Be careful not to include the mass of the cup in your measurement!**
2. Transfer the powder into the assigned "mixing cup".
3. Add water into the "mixing cup" until you have 100 mL of solution. Stir with a spoon or stick.
4. Observe and taste the solutions you have made. You can have one "designated taster" or you can pour a little into separate little cups for each group member to taste. Record in data table.
5. Repeat steps 1-5 in order to make the 0.4 M and 0.7 M solutions.
6. Clean-Up... Throw away cups that you drank out of, but save the mixing & measuring cups.

Data / Observations:

	0.1 M solution	0.4 M solution	0.7 M solution
Observations (taste (color), smell, texture)	Answers will vary	Answers will vary	Answers will vary

Calculations / Analysis:

1. Which concentration that you tasted was closest to the ideal concentration of NaCl-Aid? What was wrong with each of the other solutions that you made?

Answers will vary... Some solutions will not have enough NaCl-Aid for your taste, while others will have too much.

2. What was the solute used in this lab? What was the solvent?

Solute = NaCl-Aid powder Solvent = water

3. What is molarity?

Molarity is a measure of the number of moles of solute per liter of solution.

Questions for Discussion:... You must show the formula, all your work, and **highlight** your answer!

1. Write the formula for the 4 different ways to calculate concentrations:

- a. Molarity

$$M = \frac{\text{moles}}{\text{volume}}$$

- a. grams/Liter

$$\frac{\text{grams of solute}}{\text{volume of solution}}$$

- a. % composition

$$\frac{\text{grams of solute}}{\text{grams of solution}} \times 100$$

- a. Parts per million

$$\frac{\text{grams of solute}}{\text{grams of solution}} \times 10^6$$

Determine the following concentrations, using the 4 different methods:

****Disclaimer: 1 gram NaCl = 1 mL of 1% = 1000 mL****

1. If 1.80 moles of NaCl was dissolved in enough water to make 0.60 L of solution:

- a. Molarity?

$$M = \frac{1.8 \text{ mol}}{0.6 \text{ L}}$$

$$M = 3.000 \text{ M NaCl}$$

- a. Grams/Liter?

$$\frac{180 \text{ g NaCl}}{0.6 \text{ L}}$$

$$= 300 \text{ g/L}$$

- a. % composition?

$$\frac{180 \text{ g NaCl}}{600 \text{ g}} \times 100$$

$$= 30.0\% \text{ NaCl}$$

- a. Parts per million?

$$\frac{180 \text{ g NaCl}}{600 \text{ g}} \times 10^6$$

$$= 300,000 \text{ ppm}$$

2. If 20 grams of NaOH was dissolved in enough water to make a 250 mL NaOH solution:

- a. Molarity?

$$M = \frac{0.5 \text{ mol}}{0.25 \text{ L}}$$

$$M = 2.0 \text{ M NaOH}$$

- a. Grams/Liter?

$$\frac{20 \text{ g NaOH}}{0.25 \text{ L}}$$

$$= 80 \text{ g/L}$$

- a. % composition?

$$\frac{20 \text{ g NaOH}}{250 \text{ g}} \times 100$$

$$= 8\% \text{ NaOH}$$

- a. Parts per million?

$$\frac{20 \text{ g NaOH}}{250 \text{ g}} \times 10^6$$

$$= 80,000 \text{ ppm}$$

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Table of Contents Kool Aid Lab Answer Key

1. Understanding the eBook Kool Aid Lab Answer Key
 - The Rise of Digital Reading Kool Aid Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Kool Aid Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Kool Aid Lab Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Kool Aid Lab Answer Key
 - Personalized Recommendations
 - Kool Aid Lab Answer Key User Reviews and Ratings
 - Kool Aid Lab Answer Key and Bestseller Lists
5. Accessing Kool Aid Lab Answer Key Free and Paid eBooks
 - Kool Aid Lab Answer Key Public Domain eBooks
 - Kool Aid Lab Answer Key eBook Subscription Services
 - Kool Aid Lab Answer Key Budget-Friendly Options
6. Navigating Kool Aid Lab Answer Key eBook Formats

- ePub, PDF, MOBI, and More
- Kool Aid Lab Answer Key Compatibility with Devices
- Kool Aid Lab Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Kool Aid Lab Answer Key
 - Highlighting and Note-Taking Kool Aid Lab Answer Key
 - Interactive Elements Kool Aid Lab Answer Key
- 8. Staying Engaged with Kool Aid Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Kool Aid Lab Answer Key
- 9. Balancing eBooks and Physical Books Kool Aid Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Kool Aid Lab Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Kool Aid Lab Answer Key
 - Setting Reading Goals Kool Aid Lab Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Kool Aid Lab Answer Key
 - Fact-Checking eBook Content of Kool Aid Lab Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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