

## Chemistry Chapter 8 Exam

### Recall/Identify

Answer each question. Write the best answer below.

\_\_\_\_\_ 1. Chemical bonding

\_\_\_\_\_ 2. Ionic bonding

\_\_\_\_\_ 3. Covalent bonding

\_\_\_\_\_ 4. Electronegativity

\_\_\_\_\_ 5. Bond energy

\_\_\_\_\_ 6. Single bond

\_\_\_\_\_ 7. Multiple bonds

\_\_\_\_\_ 8. Lattice energy

\_\_\_\_\_ 9. Polyatomic ions

\_\_\_\_\_ 10. Dipole

\_\_\_\_\_ 11. Interact by oppositely charged charges that are large resulting in ionic structures

\_\_\_\_\_ 12. charge groups of negatively charged atoms

\_\_\_\_\_ 13. the energy released when ions combine with ions crystalline compounds (strongest bond)

\_\_\_\_\_ 14. ability to pull bonds

\_\_\_\_\_ 15. nonmetallic atoms that are part of molecules is called electronegativity

\_\_\_\_\_ 16. the energy required to break an chemical bond and form neutral molecules

\_\_\_\_\_ 17. chemical groups of atoms that are held together by covalent bonds

\_\_\_\_\_ 18. results from the sharing of electrons among nonmetallic atoms

\_\_\_\_\_ 19. Chemical bonding that results from the electrical attraction between positive and negative

\_\_\_\_\_ 20. mutual electrical attraction between protons and electrons in atoms, and between protons and electrons in different atoms, particularly those atoms together

### Multiple-Choice

\_\_\_\_\_ 21. The strongest bond is

- (a) ionic bond
- (b) covalent bond
- (c) metallic bond
- (d) hydrogen bond

\_\_\_\_\_ 22. In many compounds, atoms of electronegative elements form bonds with the atoms of elements of low electronegativity forming off-center dipoles

- (a) 1
- (b) 2
- (c) 3
- (d) 4

\_\_\_\_\_ 23. Which principle states that atoms tend to form compounds with the most stable electronic structure for the maximum energy level?

# Holt Mcdougal Modern Chemistry Chapter 6 Test Answers

**Yijin Wang**



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