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CHAPTER 2 – The Chemistry of Life

CHAPTER OUTLINE

2.1 Atoms Make Up All Matter

- A. Elements Are Fundamental Types of Matter
- B. Atoms Are Particles of Elements
- C. Isotopes Have Different Numbers of Neutrons

2.2 Chemical Bonds Link Atoms

- A. Electrons Determine Bonding
- B. In an Ionic Bond, One Atom Transfers Electrons to Another Atom
- C. In a Covalent Bond, Atoms Share Electrons
- D. Partial Charges on Polar Molecules Create Hydrogen Bonds

2.3 Water Is Essential to Life

- A. Water Is Cohesive and Adhesive
- B. Many Substances Dissolve in Water
- C. Water Regulates Temperature
- D. Water Expands as It Freezes
- E. Water Participates in Life's Chemical Reactions

2.4 Cells Have an Optimum pH

2.5 Cells Contain Four Major Types of Organic Molecules

- A. Large Organic Molecules Are Composed of Smaller Subunits
- B. Carbohydrates Include Simple Sugars and Polysaccharides
- C. Proteins Are Complex and Highly Versatile
- D. Nucleic Acids Store and Transmit Genetic Information
- E. Lipids Are Hydrophobic and Energy-Rich

LEARNING OUTCOMES

- 02.00.01 Explain the relationship between chemistry and biology.
- 02.01.01 Identify the most important elements in living organisms.
- 02.01.02 Describe the structure of atoms.
- 02.02.01 Compare and contrast the different types of bonds.
- 02.02.02 Differentiate between atoms and molecules.

Hoefnagels Lab Manual

Lingsheng Yao



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Table of Contents Hoefnagels Lab Manual

1. Understanding the eBook Hoefnagels Lab Manual
 - The Rise of Digital Reading Hoefnagels Lab Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Hoefnagels Lab Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hoefnagels Lab Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Hoefnagels Lab Manual
 - Personalized Recommendations
 - Hoefnagels Lab Manual User Reviews and Ratings
 - Hoefnagels Lab Manual and Bestseller Lists
5. Accessing Hoefnagels Lab Manual Free and Paid eBooks

- Hoefnagels Lab Manual Public Domain eBooks
- Hoefnagels Lab Manual eBook Subscription Services
- Hoefnagels Lab Manual Budget-Friendly Options
- 6. Navigating Hoefnagels Lab Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Hoefnagels Lab Manual Compatibility with Devices
 - Hoefnagels Lab Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hoefnagels Lab Manual
 - Highlighting and Note-Taking Hoefnagels Lab Manual
 - Interactive Elements Hoefnagels Lab Manual
- 8. Staying Engaged with Hoefnagels Lab Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hoefnagels Lab Manual
- 9. Balancing eBooks and Physical Books Hoefnagels Lab Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hoefnagels Lab Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hoefnagels Lab Manual
 - Setting Reading Goals Hoefnagels Lab Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hoefnagels Lab Manual
 - Fact-Checking eBook Content of Hoefnagels Lab Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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