

CHAPTER 21 REVIEW

Nuclear Chemistry**SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ Based on the information about the three elementary particles in the text, which has the greatest mass?
 - (a) the proton
 - (b) the neutron
 - (c) the electron
 - (d) They all have the same mass.
2. _____ The force that keeps nucleons together is
 - (a) a strong nuclear force.
 - (b) a weak nuclear force.
 - (c) an electromagnetic force.
 - (d) a gravitational force.
3. _____ The stability of a nucleus is most affected by the
 - (a) number of neutrons.
 - (b) number of protons.
 - (c) number of electrons.
 - (d) ratio of neutrons to protons.
4. _____ If an atom should form from its constituent particles,
 - (a) matter is lost and energy is taken in.
 - (b) matter is lost and energy is released.
 - (c) matter is gained and energy is taken in.
 - (d) matter is gained and energy is released.
5. _____ For atoms of a given mass number, those with greater mass defects, have
 - (a) smaller binding energies per nucleon.
 - (b) greater binding energies per nucleon.
 - (c) the same binding energies per nucleon as those with smaller mass defects.
 - (d) variable binding energies per nucleon.
6. Based on **Figure 1.1** of the text, which isotope of He, helium-3 or helium-4,
 - _____ a. has the smaller binding energy per nucleon?
 - _____ b. is more stable to nuclear changes?
7. The number of neutrons in an atom of magnesium-25 is _____.
8. Nuclides of the same element have the same _____.

Glencoe Chemistry Chapter 13 Answer

CL Gary



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(sides=2), the significance level is 0.05, and the test is to compare the difference between two means ($\mu_1 - \mu_2$) against 0

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