

Honors Final Review

Note: this is in no way guaranteed to be comprehensive. Anything from the year is fair game, but the emphasis will be on more recent material. Bolded items are most important.

Chapter 1: Introductory stuff.

- Qualitative vs quantitative measurements.
- Scientific method

Chapter 2: Atoms, elements, and mixtures.

- Matter is made of up atoms
- **Elements, compounds, and molecules: elements are only one type of atom, compounds have multiple elements, and molecules are atoms that are stuck together.**
- States of matter—solid, liquid, and gas—and their properties.
- Physical versus chemical changes.
- Mixtures: homogeneous and heterogeneous.

Chapter 3: The elements and ions.

- Dalton's atomic theory: elements are made of atoms, all atoms of a given element are identical, atoms of different elements are different, and reactions don't alter the atoms—only rearrange them.
- Structure of the atom: protons and neutrons in the nucleus, electrons around the outside.
- J.J. Thomson discovered electrons.
- Rutherford shot particles at gold foil and saw some scatter, some go through, meaning that part of the atom was hard while most was empty space.
- Element symbols: atomic number and atomic mass
- **Atomic number = number of protons, defines an element.**
- **Atomic mass = # protons + # neutrons**
- **Isotopes have same number of protons, different number of neutrons (same element, but weigh different amounts).**
- **Number of electrons changes, leading to charges. Things with a charge are called ions.**
- Positively charged atoms/groups of atoms are cations. Negatively charged ones are anions.
- Periodic table: alkali metals, alkaline earth metals, transition metals, metalloids, nonmetals, halogens, and noble gases.
- **Certain groups of elements always form ions with the same charge. See page 74.**
- Compounds containing ions conduct electricity if the ions can move.
- **Writing formulas for ionic compounds: the charges have to balance out to neutral. Use the crossing method with the charges, but reduce if there's a common factor. For example:**
 Mg^{2+} and N^{3-} for Mg_3N_2 but Mg^{2+} and O^{2-} form MgO

Chapter 4: Naming compounds.

- **Ionic compounds have a metal and at least one non-metal, and are stuck together by charges. Covalent compounds have only non-metals, and are held together by bonds.**
- Anions of a single element use the start of the element name, and end in *-ide*. Nitrogen with a negative charge becomes nitride. Oxygen with a negative charge is oxide.
- Ions can be a group of elements all stuck together by covalent bonds, that overall have a charge.
- If the group has an element attached to some number of oxygens, then it ends in *-ate*, and starts with the beginning of the name of the non-oxygen element. For example, NO_3^- is nitrate, PO_4^{3-} is phosphate.
- You can have a different number of oxygens, which have different naming schemes. The one that ends in *-ate* is the most common one.

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Rachel S Tattersall



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