

# Sulfur clock lab

## Experimental procedure

1. Measurement of the formation of visible product
2. Measurement of the initial concentration of reactants
3. Prepare solutions A, B, C, D and E as shown:

see Table 10.10.10.10

	Reactants	Reagents	Apparatus
A	1 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	Reaction setup
B	1 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	0.1 M, 0.01 M, 0.01 M
C	1 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	0.1 M, 0.01 M, 0.01 M
D	1 mL	0.1 M $\text{H}_2\text{O}_2$ , 0.01 M $\text{KSCN}$	0.1 M, 0.01 M, 0.01 M

4. Add the solution using a graduated cylinder with completely standard chemical formulae for the reaction.
5. Transfer exactly the solution into a 100 mL beaker in the reaction mixture. Transfer the solution into a 100 mL beaker with 0.1 M of  $\text{H}_2\text{O}_2$  and 0.01 M of  $\text{KSCN}$ .
6. Add exactly 100 mL of the reaction mixture at the 100 mL mark of the flask.
7. Prepare and insert the flask immediately.
8. Prepare solution B, C, D, E as shown:
  - a. Transfer 100 mL of 0.1 M  $\text{H}_2\text{O}_2$  into a beaker
  - b. Add 100 mL of 0.01 M  $\text{KSCN}$  into the beaker using a graduated cylinder
  - c. Add the 100 mL of 0.1 M  $\text{H}_2\text{O}_2$  into the beaker
  - d. Add the 100 mL of 0.01 M  $\text{KSCN}$  into the beaker
9. Transfer solution A into the reaction mixture in a 100 mL beaker. Do not stir the mixture with any efficient mixer.

# Lab Sulfur Clock Answers

**M Walker**



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