

THERMODYNAMICS OF IDEAL GASES

$$W = \frac{nR(T_2 - T_1)}{k - 1}$$

$$Q = nC_v \Delta T$$

$$\Delta U = Q + W$$

$$k = \frac{C_p}{C_v}$$

$$R = C_p - C_v$$

$$\Delta H = nC_p \Delta T$$

$$P_1 V_1^k = P_2 V_2^k$$

$$\frac{T_2}{T_1} = \left(\frac{P_2}{P_1} \right)^{\frac{k-1}{k}}$$

$$\frac{T_2}{T_1} = \left(\frac{V_1}{V_2} \right)^{k-1}$$



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Ideal Gases Section Review

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